Crystal lattice types

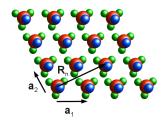
- Crystal lattice types
- The reciprocal lattice

- Crystal lattice types
- The reciprocal lattice
- Compton (inelastic) scattering

- Crystal lattice types
- The reciprocal lattice
- Compton (inelastic) scattering
- X-ray absorption

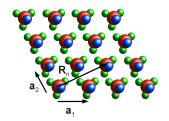
Recall that for a crystal lattice which is a periodic array of molecules

Recall that for a crystal lattice which is a periodic array of molecules



$$F^{crystal}(\mathbf{Q}) = F^{molecule}F^{lattice}$$

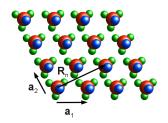
Recall that for a crystal lattice which is a periodic array of molecules



$$F^{crystal}(\mathbf{Q}) = F^{molecule}F^{lattice}$$

$$F^{crystal}(\mathbf{Q}) = \sum_{j} f_{j}(\mathbf{Q}) e^{i\mathbf{Q} \cdot \mathbf{r}_{j}} \sum_{n} e^{i\mathbf{Q} \cdot \mathbf{R}_{n}}$$

Recall that for a crystal lattice which is a periodic array of molecules

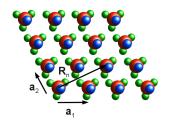


$$F^{crystal}(\mathbf{Q}) = F^{molecule}F^{lattice}$$

$$F^{crystal}(\mathbf{Q}) = \sum_{j} f_{j}(\mathbf{Q}) e^{i\mathbf{Q} \cdot \mathbf{r}_{j}} \sum_{n} e^{i\mathbf{Q} \cdot \mathbf{R}_{n}}$$

The lattice term, $\sum e^{i\mathbf{Q}\cdot\mathbf{R}_n}$, is a sum over a large number

Recall that for a crystal lattice which is a periodic array of molecules

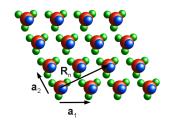


$$F^{crystal}(\mathbf{Q}) = F^{molecule}F^{lattice}$$

$$F^{crystal}(\mathbf{Q}) = \sum_{j} f_{j}(\mathbf{Q}) e^{i\mathbf{Q} \cdot \mathbf{r}_{j}} \sum_{n} e^{i\mathbf{Q} \cdot \mathbf{R}_{n}}$$

The lattice term, $\sum e^{i\mathbf{Q}\cdot\mathbf{R}_n}$, is a sum over a large number so it is always small unless $\mathbf{Q}\cdot\mathbf{R}_n=2\pi m$ where $\mathbf{R}_n=n_1\mathbf{a}_1+n_2\mathbf{a}_2+n_3\mathbf{a}_3$ is a real space lattice vector and m is an integer.

Recall that for a crystal lattice which is a periodic array of molecules



$$F^{crystal}(\mathbf{Q}) = F^{molecule}F^{lattice}$$

$$F^{crystal}(\mathbf{Q}) = \sum_{j} f_{j}(\mathbf{Q}) e^{i\mathbf{Q} \cdot \mathbf{r}_{j}} \sum_{n} e^{i\mathbf{Q} \cdot \mathbf{R}_{n}}$$

The lattice term, $\sum e^{i\mathbf{Q}\cdot\mathbf{R}_n}$, is a sum over a large number so it is always small unless $\mathbf{Q}\cdot\mathbf{R}_n=2\pi m$ where $\mathbf{R}_n=n_1\mathbf{a}_1+n_2\mathbf{a}_2+n_3\mathbf{a}_3$ is a real space lattice vector and m is an integer.

This condition is fulfilled only when \mathbf{Q} is a reciprocal lattice vector.

There are 7 possible real space lattices: triclinic,



There are 7 possible real space lattices: triclinic, monoclinic,





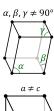
There are 7 possible real space lattices: triclinic, monoclinic, orthorhombic,







There are 7 possible real space lattices: triclinic, monoclinic, orthorhombic, tetragonal,









There are 7 possible real space lattices: triclinic, monoclinic, orthorhombic, tetragonal, hexagonal,











There are 7 possible real space lattices: triclinic, monoclinic, orthorhombic, tetragonal, hexagonal, rhombohedral,













There are 7 possible real space lattices: triclinic, monoclinic, orthorhombic, tetragonal, hexagonal, rhombohedral, cubic



















$$\mathbf{a}_1 = a\mathbf{\hat{x}}, \quad \mathbf{a}_2 = b\mathbf{\hat{y}}, \quad \mathbf{a}_3 = c\mathbf{\hat{z}}$$



$$\mathbf{a}_1 = a\mathbf{\hat{x}}, \quad \mathbf{a}_2 = b\mathbf{\hat{y}}, \quad \mathbf{a}_3 = c\mathbf{\hat{z}}$$

$$\mathbf{a}_1 \times \mathbf{a}_2 = ab\mathbf{\hat{z}}$$



$$\mathbf{a}_1 = a\hat{\mathbf{x}}, \quad \mathbf{a}_2 = b\hat{\mathbf{y}}, \quad \mathbf{a}_3 = c\hat{\mathbf{z}}$$

 $\mathbf{a}_1 \times \mathbf{a}_2 = ab\hat{\mathbf{z}}$
 $(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = ab\hat{\mathbf{z}} \cdot c\hat{\mathbf{z}}$



$$\mathbf{a}_1 = a\mathbf{\hat{x}}, \quad \mathbf{a}_2 = b\mathbf{\hat{y}}, \quad \mathbf{a}_3 = c\mathbf{\hat{z}}$$

$$\mathbf{a}_1 \times \mathbf{a}_2 = ab\mathbf{\hat{z}}$$

$$(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = ab\mathbf{\hat{z}} \cdot c\mathbf{\hat{z}}$$

$$(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = abc = V$$

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).



$$\mathbf{a}_1 = a\hat{\mathbf{x}}, \quad \mathbf{a}_2 = b\hat{\mathbf{y}}, \quad \mathbf{a}_3 = c\hat{\mathbf{z}}$$

$$\mathbf{a}_1 \times \mathbf{a}_2 = ab\hat{\mathbf{z}}$$

$$(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = ab\hat{\mathbf{z}} \cdot c\hat{\mathbf{z}}$$

$$(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = abc = V$$

A simple way of calculating the volume of the unit cell!

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).



$$\mathbf{a}_1 = a\hat{\mathbf{x}}, \quad \mathbf{a}_2 = b\hat{\mathbf{y}}, \quad \mathbf{a}_3 = c\hat{\mathbf{z}}$$

 $\mathbf{a}_1 \times \mathbf{a}_2 = ab\hat{\mathbf{z}}$
 $(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = ab\hat{\mathbf{z}} \cdot c\hat{\mathbf{z}}$
 $(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = abc = V$

A simple way of calculating the volume of the unit cell!

This unit cell is repeated infintely in 3-dimensions and thus, the location of each lattice point can be calculated relative to any arbitrary lattice point designated as the origin.

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).



$$\mathbf{a}_1 = a\hat{\mathbf{x}}, \quad \mathbf{a}_2 = b\hat{\mathbf{y}}, \quad \mathbf{a}_3 = c\hat{\mathbf{z}}$$

$$\mathbf{a}_1 \times \mathbf{a}_2 = ab\hat{\mathbf{z}}$$

$$(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = ab\hat{\mathbf{z}} \cdot c\hat{\mathbf{z}}$$

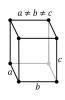
$$(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = abc = V$$

A simple way of calculating the volume of the unit cell!

This unit cell is repeated infintely in 3-dimensions and thus, the location of each lattice point can be calculated relative to any arbitrary lattice point designated as the origin.

Each lattice point is at the end of a lattice vector, \mathbf{R}_n

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).



$$\mathbf{a}_1 = a\hat{\mathbf{x}}, \quad \mathbf{a}_2 = b\hat{\mathbf{y}}, \quad \mathbf{a}_3 = c\hat{\mathbf{z}}$$

$$\mathbf{a}_1 \times \mathbf{a}_2 = ab\hat{\mathbf{z}}$$

$$(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = ab\hat{\mathbf{z}} \cdot c\hat{\mathbf{z}}$$

$$(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = abc = V$$

A simple way of calculating the volume of the unit cell!

This unit cell is repeated infintely in 3-dimensions and thus, the location of each lattice point can be calculated relative to any arbitrary lattice point designated as the origin.

Each lattice point is at the end of a lattice vector, \mathbf{R}_n

$$\mathbf{R}_n = n_1 \mathbf{a}_1 + n_2 \mathbf{a}_2 + n_3 \mathbf{a}_3$$

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).



$$\mathbf{a}_1 = a\hat{\mathbf{x}}, \quad \mathbf{a}_2 = b\hat{\mathbf{y}}, \quad \mathbf{a}_3 = c\hat{\mathbf{z}}$$

$$\mathbf{a}_1 \times \mathbf{a}_2 = ab\hat{\mathbf{z}}$$

$$(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = ab\hat{\mathbf{z}} \cdot c\hat{\mathbf{z}}$$

$$(\mathbf{a}_1 \times \mathbf{a}_2) \cdot \mathbf{a}_3 = abc = V$$

A simple way of calculating the volume of the unit cell!

This unit cell is repeated infintely in 3-dimensions and thus, the location of each lattice point can be calculated relative to any arbitrary lattice point designated as the origin.

Each lattice point is at the end of a lattice vector, \mathbf{R}_n and a crystal is made by putting a molecule at each lattice point.

$$\mathbf{R}_n = n_1 \mathbf{a}_1 + n_2 \mathbf{a}_2 + n_3 \mathbf{a}_3$$

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

$$\mathbf{a}_1^* = 2\pi \frac{\mathbf{a}_2 \times \mathbf{a}_3}{\mathbf{a}_1 \cdot (\mathbf{a}_2 \times \mathbf{a}_3)}$$

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

$$\mathbf{a}_1^* = 2\pi \frac{\mathbf{a}_2 \times \mathbf{a}_3}{\mathbf{a}_1 \cdot (\mathbf{a}_2 \times \mathbf{a}_3)}$$
$$\mathbf{a}_2^* = 2\pi \frac{\mathbf{a}_3 \times \mathbf{a}_1}{\mathbf{a}_2 \cdot (\mathbf{a}_3 \times \mathbf{a}_1)}$$

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

$$\mathbf{a}_1^* = 2\pi \frac{\mathbf{a}_2 \times \mathbf{a}_3}{\mathbf{a}_1 \cdot (\mathbf{a}_2 \times \mathbf{a}_3)}$$
$$\mathbf{a}_2^* = 2\pi \frac{\mathbf{a}_3 \times \mathbf{a}_1}{\mathbf{a}_2 \cdot (\mathbf{a}_3 \times \mathbf{a}_1)}$$
$$\mathbf{a}_3^* = 2\pi \frac{\mathbf{a}_1 \times \mathbf{a}_2}{\mathbf{a}_3 \cdot (\mathbf{a}_1 \times \mathbf{a}_2)}$$

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

$$\mathbf{a}_{1}^{*} = 2\pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot (\mathbf{a}_{2} \times \mathbf{a}_{3})} = 2\pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}$$

$$\mathbf{a}_{2}^{*} = 2\pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{\mathbf{a}_{2} \cdot (\mathbf{a}_{3} \times \mathbf{a}_{1})} = 2\pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}$$

$$\mathbf{a}_{3}^{*} = 2\pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{\mathbf{a}_{3} \cdot (\mathbf{a}_{1} \times \mathbf{a}_{2})} = 2\pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}$$

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

$$\mathbf{a}_{1}^{*} = 2\pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot (\mathbf{a}_{2} \times \mathbf{a}_{3})} = 2\pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}$$

$$\mathbf{a}_{2}^{*} = 2\pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{\mathbf{a}_{2} \cdot (\mathbf{a}_{3} \times \mathbf{a}_{1})} = 2\pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}$$

$$\mathbf{a}_{3}^{*} = 2\pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{\mathbf{a}_{3} \cdot (\mathbf{a}_{1} \times \mathbf{a}_{2})} = 2\pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}$$

In analogy to \mathbf{R}_n , we can construct an arbitrary reciprocal space lattice vector \mathbf{G}_{hkl}

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

$$\mathbf{a}_{1}^{*} = 2\pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot (\mathbf{a}_{2} \times \mathbf{a}_{3})} = 2\pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}$$

$$\mathbf{a}_{2}^{*} = 2\pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{\mathbf{a}_{2} \cdot (\mathbf{a}_{3} \times \mathbf{a}_{1})} = 2\pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}$$

$$\mathbf{a}_{3}^{*} = 2\pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{\mathbf{a}_{3} \cdot (\mathbf{a}_{1} \times \mathbf{a}_{2})} = 2\pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}$$

In analogy to \mathbf{R}_n , we can construct an arbitrary reciprocal space lattice vector \mathbf{G}_{hkl}

$$\mathbf{G}_{hkl} = h\mathbf{a}_1^* + k\mathbf{a}_2^* + l\mathbf{a}_3^*$$

Reciprocal lattice

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

$$\mathbf{a}_{1}^{*} = 2\pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot (\mathbf{a}_{2} \times \mathbf{a}_{3})} = 2\pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}$$

$$\mathbf{a}_{2}^{*} = 2\pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{\mathbf{a}_{2} \cdot (\mathbf{a}_{3} \times \mathbf{a}_{1})} = 2\pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}$$

$$\mathbf{a}_{3}^{*} = 2\pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{\mathbf{a}_{3} \cdot (\mathbf{a}_{1} \times \mathbf{a}_{2})} = 2\pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}$$

In analogy to \mathbf{R}_n , we can construct an arbitrary reciprocal space lattice vector \mathbf{G}_{hkl}

$$\mathbf{G}_{hkl} = h\mathbf{a}_1^* + k\mathbf{a}_2^* + l\mathbf{a}_3^*$$

where h, k, and I are integers called Miller indices

Because of the construction of the reciprocal lattice

 $\mathbf{G}_{hkl}\cdot\mathbf{R}_{n}$

$$\mathbf{G}_{hkl} \cdot \mathbf{R}_n = (n_1 \mathbf{a}_1 + n_2 \mathbf{a}_2 + n_3 \mathbf{a}_3) \cdot (h \mathbf{a}_1^* + k \mathbf{a}_2^* + l \mathbf{a}_3^*)$$

$$\mathbf{G}_{hkl} \cdot \mathbf{R}_{n} = (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot (h\mathbf{a}_{1}^{*} + k\mathbf{a}_{2}^{*} + l\mathbf{a}_{3}^{*})$$

$$= (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot 2\pi \left(h\frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} + k\frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} + l\frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right)$$

$$\mathbf{G}_{hkl} \cdot \mathbf{R}_{n} = (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot (h\mathbf{a}_{1}^{*} + k\mathbf{a}_{2}^{*} + l\mathbf{a}_{3}^{*})$$

$$= (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot 2\pi \left(h\frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} + k\frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} + l\frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V} \right)$$

$$= 2\pi (hn_{1} + kn_{2} + ln_{3}) = 2\pi m$$

Because of the construction of the reciprocal lattice

$$\mathbf{G}_{hkl} \cdot \mathbf{R}_{n} = (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot (h\mathbf{a}_{1}^{*} + k\mathbf{a}_{2}^{*} + l\mathbf{a}_{3}^{*})$$

$$= (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot 2\pi \left(h\frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} + k\frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} + l\frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right)$$

$$= 2\pi (hn_{1} + kn_{2} + ln_{3}) = 2\pi m$$

and therefore, the crystal scattering factor

Because of the construction of the reciprocal lattice

$$\mathbf{G}_{hkl} \cdot \mathbf{R}_{n} = (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot (h\mathbf{a}_{1}^{*} + k\mathbf{a}_{2}^{*} + I\mathbf{a}_{3}^{*})$$

$$= (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot 2\pi \left(h\frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} + k\frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} + I\frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right)$$

$$= 2\pi (hn_{1} + kn_{2} + ln_{3}) = 2\pi m$$

and therefore, the crystal scattering factor

$$\sum e^{i\mathbf{Q}\cdot\mathbf{R}_{n}}$$

Because of the construction of the reciprocal lattice

$$\mathbf{G}_{hkl} \cdot \mathbf{R}_{n} = (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot (h\mathbf{a}_{1}^{*} + k\mathbf{a}_{2}^{*} + l\mathbf{a}_{3}^{*})$$

$$= (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot 2\pi \left(h\frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} + k\frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} + l\frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right)$$

$$= 2\pi (hn_{1} + kn_{2} + ln_{3}) = 2\pi m$$

and therefore, the crystal scattering factor is non-zero only when

$$\sum e^{i{\bf Q}\cdot{\bf R}_n} \neq 0$$

Because of the construction of the reciprocal lattice

$$\mathbf{G}_{hkl} \cdot \mathbf{R}_{n} = (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot (h\mathbf{a}_{1}^{*} + k\mathbf{a}_{2}^{*} + l\mathbf{a}_{3}^{*})$$

$$= (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot 2\pi \left(h\frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} + k\frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} + l\frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right)$$

$$= 2\pi (hn_{1} + kn_{2} + ln_{3}) = 2\pi m$$

and therefore, the crystal scattering factor is non-zero only when

$$\sum e^{i\mathbf{Q}\cdot\mathbf{R}_n} \neq 0 \qquad \qquad \mathbf{Q} = \mathbf{G}_{hkl}$$

Because of the construction of the reciprocal lattice

$$\mathbf{G}_{hkl} \cdot \mathbf{R}_{n} = (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot (h\mathbf{a}_{1}^{*} + k\mathbf{a}_{2}^{*} + l\mathbf{a}_{3}^{*})$$

$$= (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot 2\pi \left(h\frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} + k\frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} + l\frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right)$$

$$= 2\pi (hn_{1} + kn_{2} + ln_{3}) = 2\pi m$$

and therefore, the crystal scattering factor is non-zero only when

$$\sum e^{i\mathbf{Q}\cdot\mathbf{R}_n} \neq 0 \qquad \qquad \mathbf{Q} = \mathbf{G}_{hkl}$$

so a significant number of molecules scatter in phase with each other

Because of the construction of the reciprocal lattice

$$\mathbf{G}_{hkl} \cdot \mathbf{R}_{n} = (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot (h\mathbf{a}_{1}^{*} + k\mathbf{a}_{2}^{*} + l\mathbf{a}_{3}^{*})$$

$$= (n_{1}\mathbf{a}_{1} + n_{2}\mathbf{a}_{2} + n_{3}\mathbf{a}_{3}) \cdot 2\pi \left(h\frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} + k\frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} + l\frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right)$$

$$= 2\pi (hn_{1} + kn_{2} + ln_{3}) = 2\pi m$$

and therefore, the crystal scattering factor is non-zero only when

$$\sum e^{i\mathbf{Q}\cdot\mathbf{R}_n} \neq 0 \qquad \qquad \mathbf{Q} = \mathbf{G}_{hkl}$$

so a significant number of molecules scatter in phase with each other

As we shall see later, this Laue condition, is equivalent to the more typically used Bragg condition for diffraction: $2d\sin\theta=n\lambda$

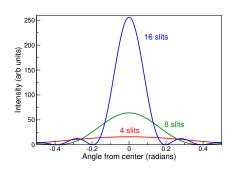
A crystal is, therefore, a diffraction grating with $\sim 10^{20}$ slits!

A crystal is, therefore, a diffraction grating with $\sim 10^{20}$ slits!

When ${\bf Q}$ is a reciprocal lattice vector, a very strong, narrow diffraction peak is seen at the detector.

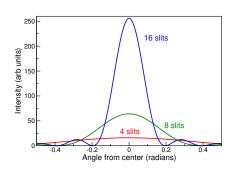
A crystal is, therefore, a diffraction grating with $\sim 10^{20}$ slits!

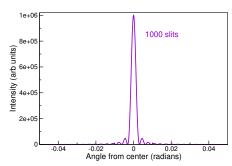
When ${\bf Q}$ is a reciprocal lattice vector, a very strong, narrow diffraction peak is seen at the detector.

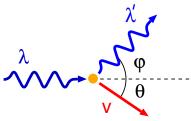


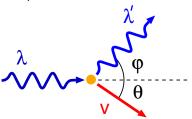
A crystal is, therefore, a diffraction grating with $\sim 10^{20}$ slits!

When ${\bf Q}$ is a reciprocal lattice vector, a very strong, narrow diffraction peak is seen at the detector.

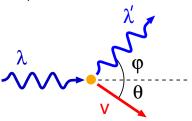




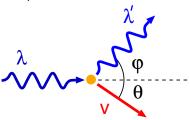




$$\mathbf{p}=\hbar\mathbf{k}=2\pi\hbar/\lambda$$

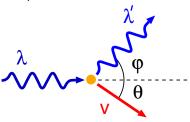


$$\mathbf{p}=\hbar\mathbf{k}=2\pi\hbar/\lambda$$
 $\mathbf{p}'=\hbar\mathbf{k}'=2\pi\hbar/\lambda'$



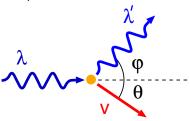
$$\mathbf{p} = \hbar \mathbf{k} = 2\pi \hbar / \lambda$$
$$\mathbf{p}' = \hbar \mathbf{k}' = 2\pi \hbar / \lambda'$$
$$|\mathbf{k}| \neq |\mathbf{k}'|$$

A photon-electron collision



$$\mathbf{p} = \hbar \mathbf{k} = 2\pi \hbar / \lambda$$
$$\mathbf{p}' = \hbar \mathbf{k}' = 2\pi \hbar / \lambda'$$
$$|\mathbf{k}| \neq |\mathbf{k}'|$$

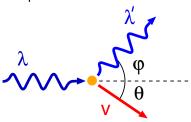
A photon-electron collision



$$\mathbf{p} = \hbar \mathbf{k} = 2\pi \hbar / \lambda$$
$$\mathbf{p}' = \hbar \mathbf{k}' = 2\pi \hbar / \lambda'$$
$$|\mathbf{k}| \neq |\mathbf{k}'|$$

$$mc^2 + \frac{hc}{\lambda} = \frac{hc}{\lambda'} + \gamma mc^2$$
 (energy)

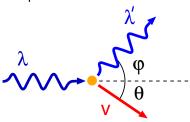
A photon-electron collision



$$\mathbf{p} = \hbar \mathbf{k} = 2\pi \hbar / \lambda$$
 $\mathbf{p}' = \hbar \mathbf{k}' = 2\pi \hbar / \lambda'$
 $|\mathbf{k}| \neq |\mathbf{k}'|$

$$mc^2 + \frac{hc}{\lambda} = \frac{hc}{\lambda'} + \gamma mc^2$$
 (energy) $\frac{h}{\lambda} = \frac{h}{\lambda'} \cos \phi + \gamma mv \cos \theta$ (x-axis)

A photon-electron collision



$$\mathbf{p}=\hbar\mathbf{k}=2\pi\hbar/\lambda$$
 $\mathbf{p}'=\hbar\mathbf{k}'=2\pi\hbar/\lambda'$
 $|\mathbf{k}|\neq |\mathbf{k}'|$

$$mc^2 + \frac{hc}{\lambda} = \frac{hc}{\lambda'} + \gamma mc^2$$
 (energy)
 $\frac{h}{\lambda} = \frac{h}{\lambda'} \cos \phi + \gamma mv \cos \theta$ (x-axis)
 $0 = \frac{h}{\lambda'} \sin \phi + \gamma mv \sin \theta$ (y-axis)

squaring the momentum equations

squaring the momentum equations

$$\left(\frac{h}{\lambda} - \frac{h}{\lambda'}\cos\phi\right)^2 = \gamma^2 m^2 v^2 \cos^2\theta$$

squaring the momentum equations

$$\left(\frac{h}{\lambda} - \frac{h}{\lambda'}\cos\phi\right)^2 = \gamma^2 m^2 v^2 \cos^2\theta$$
$$\left(-\frac{h}{\lambda'}\sin\phi\right)^2 = \gamma^2 m^2 v^2 \sin^2\theta$$

squaring the momentum equations

$$\left(\frac{h}{\lambda} - \frac{h}{\lambda'}\cos\phi\right)^2 = \gamma^2 m^2 v^2 \cos^2\theta$$
$$\left(-\frac{h}{\lambda'}\sin\phi\right)^2 = \gamma^2 m^2 v^2 \sin^2\theta$$

now add them together,

$$\gamma^2 m^2 v^2 \left(\sin^2 \theta + \cos^2 \theta \right) = \left(\frac{h}{\lambda} - \frac{h}{\lambda'} \cos \phi \right)^2 + \left(-\frac{h}{\lambda'} \sin \phi \right)^2$$

squaring the momentum equations

$$\left(\frac{h}{\lambda} - \frac{h}{\lambda'}\cos\phi\right)^2 = \gamma^2 m^2 v^2 \cos^2\theta$$
$$\left(-\frac{h}{\lambda'}\sin\phi\right)^2 = \gamma^2 m^2 v^2 \sin^2\theta$$

now add them together, substitute $\sin^2 \theta + \cos^2 \theta = 1$, expand the squares,

$$\gamma^2 m^2 v^2 \left(\sin^2 \theta + \cos^2 \theta \right) = \left(\frac{h}{\lambda} - \frac{h}{\lambda'} \cos \phi \right)^2 + \left(-\frac{h}{\lambda'} \sin \phi \right)^2$$
$$\gamma^2 m^2 v^2 = \frac{h^2}{\lambda^2} - \frac{2h^2}{\lambda \lambda'} \cos \phi + \frac{h^2}{\lambda'^2} \sin^2 \phi + \frac{h^2}{\lambda'^2} \cos^2 \phi$$

squaring the momentum equations

$$\left(\frac{h}{\lambda} - \frac{h}{\lambda'}\cos\phi\right)^2 = \gamma^2 m^2 v^2 \cos^2\theta$$
$$\left(-\frac{h}{\lambda'}\sin\phi\right)^2 = \gamma^2 m^2 v^2 \sin^2\theta$$

now add them together, substitute $\sin^2\theta+\cos^2\theta=1$, expand the squares, and $\sin^2\phi+\cos^2\phi=1$, then rearrange

$$\gamma^2 m^2 v^2 \left(\sin^2 \theta + \cos^2 \theta \right) = \left(\frac{h}{\lambda} - \frac{h}{\lambda'} \cos \phi \right)^2 + \left(-\frac{h}{\lambda'} \sin \phi \right)^2$$
$$\gamma^2 m^2 v^2 = \frac{h^2}{\lambda^2} - \frac{2h^2}{\lambda \lambda'} \cos \phi + \frac{h^2}{\lambda'^2} \sin^2 \phi + \frac{h^2}{\lambda'^2} \cos^2 \phi$$
$$\frac{m^2 v^2}{1 - \beta^2} = \frac{h^2}{\lambda^2} - \frac{2h^2}{\lambda \lambda'} \cos \phi + \frac{h^2}{\lambda'^2}$$

squaring the momentum equations

$$\left(\frac{h}{\lambda} - \frac{h}{\lambda'}\cos\phi\right)^2 = \gamma^2 m^2 v^2 \cos^2\theta$$
$$\left(-\frac{h}{\lambda'}\sin\phi\right)^2 = \gamma^2 m^2 v^2 \sin^2\theta$$

now add them together, substitute $\sin^2\theta+\cos^2\theta=1$, expand the squares, and $\sin^2\phi+\cos^2\phi=1$, then rearrange and substitute $v=\beta c$

$$\gamma^2 m^2 v^2 \left(\sin^2 \theta + \cos^2 \theta \right) = \left(\frac{h}{\lambda} - \frac{h}{\lambda'} \cos \phi \right)^2 + \left(-\frac{h}{\lambda'} \sin \phi \right)^2$$
$$\gamma^2 m^2 v^2 = \frac{h^2}{\lambda^2} - \frac{2h^2}{\lambda \lambda'} \cos \phi + \frac{h^2}{\lambda'^2} \sin^2 \phi + \frac{h^2}{\lambda'^2} \cos^2 \phi$$
$$\frac{m^2 c^2 \beta^2}{1 - \beta^2} = \frac{m^2 v^2}{1 - \beta^2} = \frac{h^2}{\lambda^2} - \frac{2h^2}{\lambda \lambda'} \cos \phi + \frac{h^2}{\lambda'^2}$$

Now take the energy equation and square it,

$$\left(mc^2 + \frac{hc}{\lambda} - \frac{hc}{\lambda'}\right)^2 = \gamma^2 m^2 c^4 = \frac{m^2 c^4}{1 - \beta^2}$$

Now take the energy equation and square it, then solve it for β^2

$$\left(mc^{2} + \frac{hc}{\lambda} - \frac{hc}{\lambda'}\right)^{2} = \gamma^{2}m^{2}c^{4} = \frac{m^{2}c^{4}}{1 - \beta^{2}}$$

$$\beta^2 = 1 - \frac{m^2 c^4}{\left(mc^2 + \frac{hc}{\lambda} - \frac{hc}{\lambda'}\right)^2}$$

Now take the energy equation and square it, then solve it for β^2 which is substituted into the equation from the momentum.

$$\left(mc^{2} + \frac{hc}{\lambda} - \frac{hc}{\lambda'}\right)^{2} = \gamma^{2}m^{2}c^{4} = \frac{m^{2}c^{4}}{1 - \beta^{2}}$$

$$\beta^2 = 1 - \frac{m^2 c^4}{\left(mc^2 + \frac{hc}{\lambda} - \frac{hc}{\lambda'}\right)^2}$$

$$\frac{h^2}{\lambda^2} + \frac{h^2}{\lambda'^2} - \frac{2h^2}{\lambda \lambda'} \cos \phi = \frac{m^2 c^2 \beta^2}{1 - \beta^2}$$

Now take the energy equation and square it, then solve it for β^2 which is substituted into the equation from the momentum.

$$\left(mc^2 + \frac{hc}{\lambda} - \frac{hc}{\lambda'}\right)^2 = \gamma^2 m^2 c^4 = \frac{m^2 c^4}{1 - \beta^2}$$

$$\beta^2 = 1 - \frac{m^2 c^4}{\left(mc^2 + \frac{hc}{\lambda} - \frac{hc}{\lambda'}\right)^2}$$

$$\frac{h^2}{\lambda^2} + \frac{h^2}{\lambda'^2} - \frac{2h^2}{\lambda \lambda'} \cos \phi = \frac{m^2 c^2 \beta^2}{1 - \beta^2}$$
$$= \frac{1}{c^2} \left(mc^2 + \frac{hc}{\lambda} - \frac{hc}{\lambda'} \right)^2 - m^2 c^2$$

$$\frac{h^2}{\lambda^2} + \frac{h^2}{\lambda'^2} - \frac{2h^2}{\lambda\lambda'}\cos\phi = \left(mc + \frac{h}{\lambda} - \frac{h}{\lambda'}\right)^2 - m^2c^2$$

After expansion,

$$\frac{h^{2}}{\lambda^{2}} + \frac{h^{2}}{\lambda'^{2}} - \frac{2h^{2}}{\lambda \lambda'} \cos \phi = \left(mc + \frac{h}{\lambda} - \frac{h}{\lambda'}\right)^{2} - m^{2}c^{2}$$

$$= m^{2}c^{2} + \frac{h^{2}}{\lambda^{2}} + \frac{h^{2}}{\lambda'^{2}} - \frac{2mch}{\lambda} - \frac{2mch}{\lambda'} + \frac{2h^{2}}{\lambda \lambda'} - m^{2}c^{2}$$

$$\frac{h^2}{\lambda^2} + \frac{h^2}{\lambda'^2} - \frac{2h^2}{\lambda \lambda'} \cos \phi = \left(mc + \frac{h}{\lambda} - \frac{h}{\lambda'}\right)^2 - m^2 c^2$$

$$= m^2 c^2 + \frac{h^2}{\lambda^2} + \frac{h^2}{\lambda'^2} - \frac{2mch}{\lambda} - \frac{2mch}{\lambda'} + \frac{2h^2}{\lambda \lambda'} - m^2 c^2$$

$$= 2m \left(\frac{hc}{\lambda} - \frac{hc}{\lambda'}\right) + \frac{h^2}{\lambda^2} + \frac{h^2}{\lambda'^2} - \frac{2h^2}{\lambda \lambda'}$$

$$\frac{h^{2}}{\sqrt{\lambda^{2}}} + \frac{h^{2}}{\sqrt{\lambda^{\prime 2}}} - \frac{2h^{2}}{\lambda \lambda^{\prime}} \cos \phi = \left(mc + \frac{h}{\lambda} - \frac{h}{\lambda^{\prime}}\right)^{2} - m^{2}c^{2}$$

$$= m^{2}c^{2} + \frac{h^{2}}{\lambda^{2}} + \frac{h^{2}}{\lambda^{\prime 2}} - \frac{2mch}{\lambda} - \frac{2mch}{\lambda^{\prime}} + \frac{2h^{2}}{\lambda \lambda^{\prime}} - m^{2}c^{2}$$

$$= 2m\left(\frac{hc}{\lambda} - \frac{hc}{\lambda^{\prime}}\right) + \frac{h^{2}}{\lambda^{2}} + \frac{h^{2}}{\lambda^{\prime 2}} - \frac{2h^{2}}{\lambda \lambda^{\prime}}$$

After expansion, cancellation, and rearrangement, we obtain

$$\frac{h^2/\sqrt{h^2} + \frac{h^2/\sqrt{h^2}}{\lambda^2} - \frac{2h^2}{\lambda \lambda'} \cos \phi}{\ln c} = \left(mc + \frac{h}{\lambda} - \frac{h}{\lambda'}\right)^2 - m^2 c^2$$

$$= m^2 c^2 + \frac{h^2}{\lambda^2} + \frac{h^2}{\lambda'^2} - \frac{2mch}{\lambda} - \frac{2mch}{\lambda'} + \frac{2h^2}{\lambda \lambda'} - m^2 c^2$$

$$= 2m \left(\frac{hc}{\lambda} - \frac{hc}{\lambda'}\right) + \frac{h^2/\sqrt{h^2}}{\lambda^2} + \frac{h^2/\sqrt{h^2}}{\lambda'^2} - \frac{2h^2}{\lambda \lambda'}$$

$$\frac{2h^2}{\lambda\lambda'}(1-\cos\phi)=2m\left(\frac{hc}{\lambda}-\frac{hc}{\lambda'}\right)$$

$$\frac{h^2/\sqrt{h^2} + \frac{h^2/\sqrt{h^2}}{\lambda^2} - \frac{2h^2}{\lambda \lambda^2} \cos \phi}{\ln c} = \left(mc + \frac{h}{\lambda} - \frac{h}{\lambda^2}\right)^2 - m^2 c^2$$

$$= m^2 c^2 + \frac{h^2}{\lambda^2} + \frac{h^2}{\lambda^2} - \frac{2mch}{\lambda} - \frac{2mch}{\lambda^2} + \frac{2h^2}{\lambda \lambda^2} - m^2 c^2$$

$$= 2m \left(\frac{hc}{\lambda} - \frac{hc}{\lambda^2}\right) + \frac{h^2/\sqrt{h^2}}{\lambda^2} + \frac{h^2/\sqrt{h^2}}{\lambda^2} - \frac{2h^2}{\lambda \lambda^2}$$

$$\frac{2h^2}{\lambda \lambda'} \left(1 - \cos \phi \right) = 2m \left(\frac{hc}{\lambda} - \frac{hc}{\lambda'} \right) = 2mhc \left(\frac{\lambda' - \lambda}{\lambda \lambda'} \right)$$

$$\frac{h^2/\sqrt{h^2} + \frac{h^2/\sqrt{h^2}}{\lambda^2} - \frac{2h^2}{\lambda^2} \cos \phi}{\ln c} = \left(mc + \frac{h}{\lambda} - \frac{h}{\lambda^2}\right)^2 - m^2c^2$$

$$= m^2c^2 + \frac{h^2}{\lambda^2} + \frac{h^2}{\lambda^2} - \frac{2mch}{\lambda} - \frac{2mch}{\lambda^2} + \frac{2h^2}{\lambda\lambda^2} - m^2c^2$$

$$= 2m\left(\frac{hc}{\lambda} - \frac{hc}{\lambda^2}\right) + \frac{h^2/\sqrt{h^2}}{\lambda^2} + \frac{h^2/\sqrt{h^2}}{\lambda^2} - \frac{2h^2}{\lambda\lambda^2}$$

$$\frac{2h^2}{\lambda\lambda'}\left(1-\cos\phi\right)=2m\left(\frac{hc}{\lambda}-\frac{hc}{\lambda'}\right)=2mhc\left(\frac{\lambda'-\lambda}{\lambda\lambda'}\right)=\frac{2mhc\Delta\lambda}{\lambda\lambda'}$$

$$\frac{h^2/\sqrt{h^2} + \frac{h^2/\sqrt{h^2}}{\lambda^2} - \frac{2h^2}{\lambda^2} \cos \phi}{\ln c} = \left(mc + \frac{h}{\lambda} - \frac{h}{\lambda^2}\right)^2 - m^2 c^2$$

$$= m^2 c^2 + \frac{h^2}{\lambda^2} + \frac{h^2}{\lambda^2} - \frac{2mch}{\lambda} - \frac{2mch}{\lambda^2} + \frac{2h^2}{\lambda\lambda^2} - m^2 c^2$$

$$= 2m\left(\frac{hc}{\lambda} - \frac{hc}{\lambda^2}\right) + \frac{h^2/\sqrt{h^2}}{\lambda^2} + \frac{h^2/\sqrt{h^2}}{\lambda^2} - \frac{2h^2}{\lambda\lambda^2}$$

$$\frac{2h^2}{\lambda \lambda'} (1 - \cos \phi) = 2m \left(\frac{hc}{\lambda} - \frac{hc}{\lambda'} \right) = 2mhc \left(\frac{\lambda' - \lambda}{\lambda \lambda'} \right) = \frac{2mhc\Delta\lambda}{\lambda \lambda'}$$

$$\Delta\lambda = \frac{h}{mc} \left(1 - \cos \phi \right)$$

Thus, for an electron

Thus, for an electron

$$\lambda_c = \hbar/mc = 3.86 \times 10^{-3} \text{Å}$$

Thus, for an electron

$$\lambda_c = \hbar/mc = 3.86 \times 10^{-3} \text{Å}$$
 $r_0 = \frac{e^2}{4\pi\epsilon_0 mc^2} = 2.82 \times 10^{-5} \text{Å}$

Thus, for an electron

Comparing the two scattering lengths:

Comparing the two scattering lengths:
$$r_0/\lambda_C = 1/137$$

$$\lambda_c = \hbar/mc = 3.86 \times 10^{-3} \text{Å}$$

$$r_0 = \frac{e^2}{4\pi\epsilon_0 mc^2} = 2.82 \times 10^{-5} \text{Å}$$

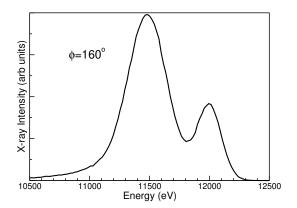
Thus, for an electron

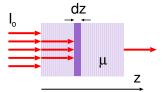
Comparing the two scattering lengths:

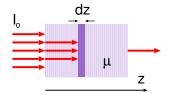
Comparing the two scattering lengths:
$$r_0/\lambda_C = 1/137$$

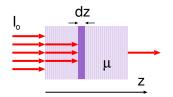
$$\lambda_c = \hbar/mc = 3.86 \times 10^{-3} \text{Å}$$

$$r_0 = \frac{e^2}{4\pi\epsilon_0 mc^2} = 2.82 \times 10^{-5} \text{Å}$$

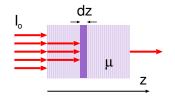




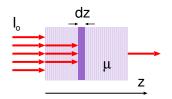




$$dI = -I(z)\mu dz$$



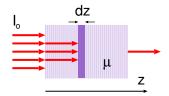
$$dI = -I(z)\mu dz \longrightarrow \frac{dI}{I(z)} = -\mu dz$$



For absorption coefficient μ and thickness dz the x-ray intensity is attenuated as

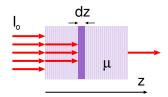
$$dI = -I(z)\mu dz \longrightarrow \frac{dI}{I(z)} = -\mu dz$$

integrating both sides



integrating both sides

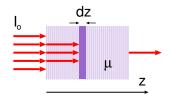
$$dI = -I(z)\mu dz \longrightarrow \frac{dI}{I(z)} = -\mu dz$$
$$\int \frac{dI}{I(z)} = -\int \mu dz$$



integrating both sides

$$dI = -I(z)\mu dz \longrightarrow \frac{dI}{I(z)} = -\mu dz$$

$$\int \frac{dI}{I(z)} = -\int \mu dz \longrightarrow \ln(I) = -\mu z + C$$

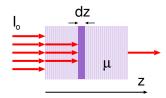


integrating both sides

and taking the anti-log

$$dI = -I(z)\mu dz \longrightarrow \frac{dI}{I(z)} = -\mu dz$$

$$\int \frac{dI}{I(z)} = -\int \mu dz \longrightarrow \ln(I) = -\mu z + C$$



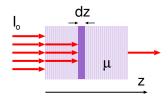
integrating both sides

and taking the anti-log

$$dI = -I(z)\mu dz \longrightarrow \frac{dI}{I(z)} = -\mu dz$$

$$\int \frac{dI}{I(z)} = -\int \mu dz \longrightarrow \ln(I) = -\mu z + C$$

$$I = e^{C}e^{-\mu z} = Ae^{-\mu z}$$



integrating both sides

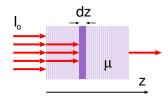
and taking the anti-log

if the intensity at z = 0 is I_0 , then

$$dI = -I(z)\mu dz \longrightarrow \frac{dI}{I(z)} = -\mu dz$$

$$\int \frac{dI}{I(z)} = -\int \mu dz \longrightarrow \ln(I) = -\mu z + C$$

$$I = e^C e^{-\mu z} = A e^{-\mu z}$$



integrating both sides

and taking the anti-log

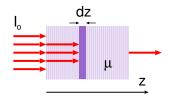
if the intensity at z = 0 is I_0 , then

$$dI = -I(z)\mu dz \longrightarrow \frac{dI}{I(z)} = -\mu dz$$

$$\int \frac{dI}{I(z)} = -\int \mu dz \longrightarrow \ln(I) = -\mu z + C$$

$$I = e^C e^{-\mu z} = A e^{-\mu z}$$

$$I = I_0 e^{-\mu z}$$



integrating both sides

and taking the anti-log

if the intensity at z = 0 is I_0 , then

For absorption coefficient μ and thickness dz the x-ray intensity is attenuated as

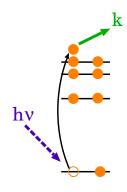
$$dI = -I(z)\mu dz \longrightarrow \frac{dI}{I(z)} = -\mu dz$$

$$\int \frac{dI}{I(z)} = -\int \mu dz \longrightarrow \ln(I) = -\mu z + C$$

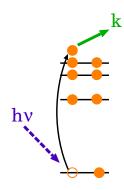
$$I = e^{C}e^{-\mu z} = Ae^{-\mu z}$$

$$I = I_0 e^{-\mu z}$$

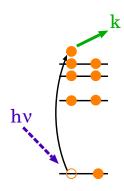
This is just Beer's law with an absorption coefficient which depends on x-ray parameters.



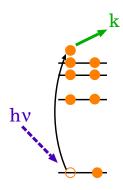
X-ray is absorbed by an atom



- X-ray is absorbed by an atom
- Energy is transferred to a core electron



- X-ray is absorbed by an atom
- Energy is transferred to a core electron
- Electron escapes atomic potential into the continuum



- X-ray is absorbed by an atom
- Energy is transferred to a core electron
- Electron escapes atomic potential into the continuum
- Ion remains with a core-hole

An ion with a core-hole is quite unstable ($\approx 10^{-15} \text{s})$

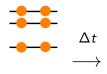






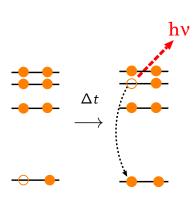
An ion with a core-hole is quite unstable ($\approx 10^{-15} s$)

 After a short time a higher level electron will drop down in energy to fill the core hole



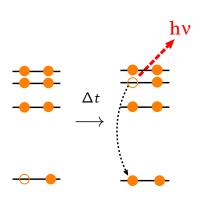


An ion with a core-hole is quite unstable ($\approx 10^{-15} s$)



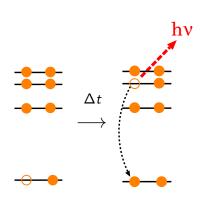
- After a short time a higher level electron will drop down in energy to fill the core hole
- Energy is liberated in the form of a fluorescence photon

An ion with a core-hole is quite unstable ($\approx 10^{-15} \mathrm{s}$)



- After a short time a higher level electron will drop down in energy to fill the core hole
- Energy is liberated in the form of a fluorescence photon
- This leaves a second hole (not core) which is then filled from an even higher shell

An ion with a core-hole is quite unstable ($\approx 10^{-15}$ s)



- After a short time a higher level electron will drop down in energy to fill the core hole
- Energy is liberated in the form of a fluorescence photon
- This leaves a second hole (not core) which is then filled from an even higher shell
- The result is a cascade of fluorescence photons which are characteristic of the absorbing atom

Auger emission

While fluorescence is the most probable method of core-hole relaxation there are other possible mechanisms







Auger emission

While fluorescence is the most probable method of core-hole relaxation there are other possible mechanisms

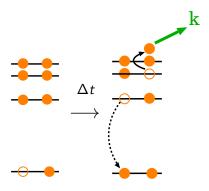


 In the Auger process, a higher level electron will drop down in energy to fill the core hole



Auger emission

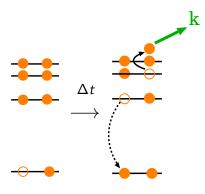
While fluorescence is the most probable method of core-hole relaxation there are other possible mechanisms



- In the Auger process, a higher level electron will drop down in energy to fill the core hole
- The energy liberated causes the secondary emission of an electron

Auger emission

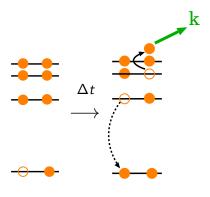
While fluorescence is the most probable method of core-hole relaxation there are other possible mechanisms



- In the Auger process, a higher level electron will drop down in energy to fill the core hole
- The energy liberated causes the secondary emission of an electron
- This leaves two holes which then filled from higher shells

Auger emission

While fluorescence is the most probable method of core-hole relaxation there are other possible mechanisms



- In the Auger process, a higher level electron will drop down in energy to fill the core hole
- The energy liberated causes the secondary emission of an electron
- This leaves two holes which then filled from higher shells
- So that the secondary electron is accompanied by fluorescence emissions at lower energies

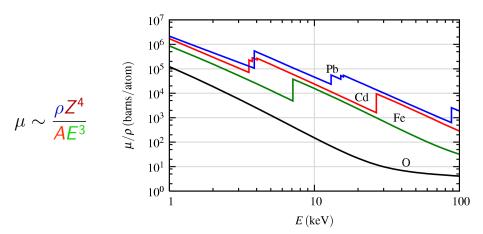


$$\mu \sim \frac{1}{E^3}$$

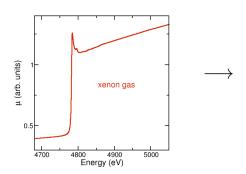
$$\mu \sim \frac{Z^4}{E^3}$$

$$\mu \sim \frac{\rho Z^4}{E^3}$$

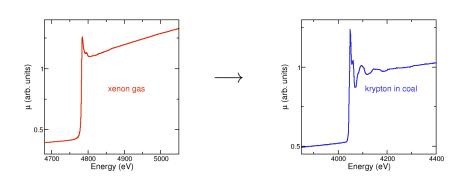
$$\mu \sim \frac{\rho Z^4}{AE^3}$$



Isolated gas atoms show a sharp jump and a smooth curve



Isolated gas atoms show a sharp jump and a smooth curve Atoms in a solid or liquid show fine structure after the absorption edge called XANES and EXAFS



For an elemental material, the absorption coefficient, μ , is simply the product of the atomic density, ρ_a , times the atomic absorption cross-section, σ_a .

For an elemental material, the absorption coefficient, μ , is simply the product of the atomic density, ρ_a , times the atomic absorption cross-section, σ_a .

$$\mu = \rho_{\mathsf{a}} \sigma_{\mathsf{a}}$$

For an elemental material, the absorption coefficient, μ , is simply the product of the atomic density, ρ_a , times the atomic absorption cross-section, σ_a .

rewriting in terms of the mass density, ρ_m , the atomic mass, M_a , and Avogadro's number, N_A , the absorption coefficient becomes

$$\mu = \rho_{\mathsf{a}} \sigma_{\mathsf{a}}$$

For an elemental material, the absorption coefficient, μ , is simply the product of the atomic density, ρ_a , times the atomic absorption cross-section, σ_a .

rewriting in terms of the mass density, ρ_m , the atomic mass, M_a , and Avogadro's number, N_A , the absorption coefficient becomes

$$\mu = \rho_{\mathsf{a}} \sigma_{\mathsf{a}} = \left(\frac{\rho_{\mathsf{m}} \mathsf{N}_{\mathsf{A}}}{\mathsf{M}_{\mathsf{a}}}\right) \sigma_{\mathsf{a}}$$

For an elemental material, the absorption coefficient, μ , is simply the product of the atomic density, ρ_a , times the atomic absorption cross-section, σ_a .

rewriting in terms of the mass density, ρ_m , the atomic mass, M_a , and Avogadro's number, N_A , the absorption coefficient becomes

$$\mu = \rho_{\mathsf{a}} \sigma_{\mathsf{a}} = \left(\frac{\rho_{\mathsf{m}} \mathsf{N}_{\mathsf{A}}}{\mathsf{M}_{\mathsf{a}}}\right) \sigma_{\mathsf{a}}$$

if the absorber is made up of a number of different atoms, this calculation can be generalized

For an elemental material, the absorption coefficient, μ , is simply the product of the atomic density, ρ_a , times the atomic absorption cross-section, σ_a .

rewriting in terms of the mass density, ρ_m , the atomic mass, M_a , and Avogadro's number, N_A , the absorption coefficient becomes

$$\mu = \rho_{\mathsf{a}} \sigma_{\mathsf{a}} = \left(\frac{\rho_{\mathsf{m}} \mathsf{N}_{\mathsf{A}}}{\mathsf{M}_{\mathsf{a}}}\right) \sigma_{\mathsf{a}}$$

if the absorber is made up of a number of different atoms, this calculation can be generalized

$$\mu = \sum_{j} \rho_{j} \sigma_{aj}$$

For an elemental material, the absorption coefficient, μ , is simply the product of the atomic density, ρ_a , times the atomic absorption cross-section, σ_a .

rewriting in terms of the mass density, ρ_m , the atomic mass, M_a , and Avogadro's number, N_A , the absorption coefficient becomes

$$\mu = \rho_{\mathsf{a}} \sigma_{\mathsf{a}} = \left(\frac{\rho_{\mathsf{m}} \mathsf{N}_{\mathsf{A}}}{\mathsf{M}_{\mathsf{a}}}\right) \sigma_{\mathsf{a}}$$

if the absorber is made up of a number of different atoms, this calculation can be generalized

$$\mu = \sum_{i} \rho_{j} \sigma_{aj}$$

where ρ_j and σ_{aj} are the atomic density and atomic absorption cross-section of each component

 $\mu[{\rm cm}^-1]$ is the linear absorption coefficient. It is useful in practice to define the mass absorption coefficient, $\mu_m[{\rm cm}^2/{\rm g}]$

 μ [cm⁻1] is the linear absorption coefficient. It is useful in practice to define the mass ab- $\mu_m = \mu/\rho$ sorption coefficient, $\mu_m[\text{cm}^2/\text{g}]$

$$\mu_{\rm m} = \mu/\rho$$

 μ [cm⁻1] is the linear absorption coefficient. It is useful in practice to define the mass ab- $\mu_m = \mu/\rho$ sorption coefficient, $\mu_m[\text{cm}^2/\text{g}]$

$$\mu_{\rm m}=\mu/\rho$$

Beer's Law now becomes

 $\mu[{\rm cm}^-1]$ is the linear absorption coefficient. It is useful in practice to define the mass absorption coefficient, $\mu_m[{\rm cm}^2/{\rm g}]$

Beer's Law now becomes

$$\mu_{\rm m}=\mu/\rho$$

$$I = I_0 e^{-\mu z} = I_0 e^{-\mu_m \rho z}$$

 μ [cm⁻1] is the linear absorption coefficient. It is useful in practice to define the mass ab- $\mu_m = \mu/\rho$ sorption coefficient, $\mu_m[\text{cm}^2/\text{g}]$

$$\mu_{\rm m} = \mu/\rho$$

Beer's Law now becomes

$$I = I_0 e^{-\mu z} = I_0 e^{-\mu_m \rho z}$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area A

 $\mu[{\rm cm}^-1]$ is the linear absorption coefficient. It is useful in practice to define the mass absorption coefficient, $\mu_m[{\rm cm}^2/{\rm g}]$

$$\mu_{\rm m} = \mu/\rho$$

Beer's Law now becomes

$$I = I_0 e^{-\mu z} = I_0 e^{-\mu_m \rho z}$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area A

If the compound is made up of x_j atoms with atomic mass M_j and has a molecular mass M_c and density ρ_c , we can write:

 $\mu[{\rm cm}^-1]$ is the linear absorption coefficient. It is useful in practice to define the mass absorption coefficient, $\mu_m[{\rm cm}^2/{\rm g}]$

$$\mu_{\rm m}=\mu/\rho$$

Beer's Law now becomes

$$I = I_0 e^{-\mu z} = I_0 e^{-\mu_m \rho z}$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area A

If the compound is made up of x_j atoms with atomic mass M_j and has a molecular mass M_c and density ρ_c , we can write:

$$M_c = \sum_j x_j M_j$$

 $\mu[{\rm cm}^-1]$ is the linear absorption coefficient. It is useful in practice to define the mass absorption coefficient, $\mu_m[{\rm cm}^2/{\rm g}]$

$$\mu_{\rm m}=\mu/\rho$$

Beer's Law now becomes

$$I = I_0 e^{-\mu z} = I_0 e^{-\mu_m \rho z}$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area A

If the compound is made up of x_j atoms with atomic mass M_j and has a molecular mass M_c and density ρ_c , we can write:

$$M_c = \sum_j x_j M_j$$
 $\mu_m = (N_A/M_c) \sum_j x_j \sigma_{aj}$

 $\mu [{\rm cm}^-1]$ is the linear absorption coefficient. It is useful in practice to define the mass absorption coefficient, $\mu_m [{\rm cm}^2/{\rm g}]$

$$\mu_{\rm m}=\mu/\rho$$

Beer's Law now becomes

$$I = I_0 e^{-\mu z} = I_0 e^{-\mu_m \rho z}$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area A

If the compound is made up of x_j atoms with atomic mass M_j and has a molecular mass M_c and density ρ_c , we can write:

$$M_c = \sum_j x_j M_j$$
 $\mu_m = (N_A/M_c) \sum_j x_j \sigma_{aj}$

The "thickness" of a mass m of the compound, distributed over an area A is then:

 $\mu [{\rm cm}^-1]$ is the linear absorption coefficient. It is useful in practice to define the mass absorption coefficient, $\mu_m [{\rm cm}^2/{\rm g}]$

$$\mu_{\rm m}=\mu/\rho$$

Beer's Law now becomes

$$I = I_0 e^{-\mu z} = I_0 e^{-\mu_m \rho z}$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area A

If the compound is made up of x_j atoms with atomic mass M_j and has a molecular mass M_c and density ρ_c , we can write:

 $M_{c} = \sum_{j} x_{j} M_{j}$ $\mu_{m} = (N_{A}/M_{c}) \sum_{j} x_{j} \sigma_{aj}$ $z = \frac{m}{\rho_{c} A}$

The "thickness" of a mass m of the compound, distributed over an area A is then:

 $\mu[{\rm cm}^-1]$ is the linear absorption coefficient. It is useful in practice to define the mass absorption coefficient, $\mu_m[{\rm cm}^2/{\rm g}]$

$$\mu_{\rm m}=\mu/\rho$$

Beer's Law now becomes

$$I = I_0 e^{-\mu z} = I_0 e^{-\mu_m \rho z}$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area A

If the compound is made up of x_j atoms with atomic mass M_j and has a molecular mass M_c and density ρ_c , we can write:

 $M_c = \sum_j x_j M_j$ $\mu_m = (N_A/M_c) \sum_j x_j \sigma_{aj}$ $z = \frac{m}{\rho_c A}$

The "thickness" of a mass m of the compound, distributed over an area A is then:

This leads to an absorption per unit mass of μ_m/A and Beer's law becomes

 $\mu[{\rm cm}^-1]$ is the linear absorption coefficient. It is useful in practice to define the mass absorption coefficient, $\mu_m[{\rm cm}^2/{\rm g}]$

$$\mu_{\rm m}=\mu/\rho$$

Beer's Law now becomes

$$I = I_0 e^{-\mu z} = I_0 e^{-\mu_m \rho z}$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area A

If the compound is made up of x_j atoms with atomic mass M_j and has a molecular mass M_c and density ρ_c , we can write:

 $M_{c} = \sum_{j} x_{j} M_{j}$ $\mu_{m} = (N_{A}/M_{c}) \sum_{j} x_{j} \sigma_{aj}$ $z = \frac{m}{\rho_{c} A}$ $I = I_{0} e^{-(\mu_{m}/A)m}$

The "thickness" of a mass m of the compound, distributed over an area A is then:

This leads to an absorption per unit mass of μ_m/A and Beer's law becomes

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_i x_i \sigma_{ai}$$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_i x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_i \frac{M_j}{N_A} x_j \sigma_j$$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_i x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_i \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_i M_j x_j \sigma_j$$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_i x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_i \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_i M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of $\mathrm{Fe_2O_3}$ are

$$ho = 5.24\,\mathrm{g/cm^2}$$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of $\operatorname{Fe_2O_3}$ are

$$ho = 5.24 \, {
m g/cm^2}$$
 $M_{Fe} = 55.895 \, {
m g/mol}$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j=N_A\sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

$$ho = 5.24 \, {
m g/cm}^2$$

 $M_{Fe} = 55.895 \, {
m g/mol}$
 $M_O = 16.000 \, {
m g/mol}$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of $\operatorname{Fe_2O_3}$ are

$$ho=5.24\,\mathrm{g/cm^2}$$
 $M_{Fe}=55.895\,\mathrm{g/mol}$
 $M_O=16.000\,\mathrm{g/mol}$
 $M_C=159.69\,\mathrm{g/mol}$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

begin by finding tabulated values of the cross-section for the elements Fe and O at $5\ \text{keV}$

$$\rho = 5.24\,\mathrm{g/cm^2}$$

$$M_{Fe} = 55.895 \,\mathrm{g/mol}$$

$$M_O = 16.000\,\mathrm{g/mol}$$

$$M_c=159.69\,\mathrm{g/mol}$$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j=N_A\sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

begin by finding tabulated values of the cross-section for the elements Fe and O at $5\ \text{keV}$

$$\rho=5.24\,\mathrm{g/cm^2}$$

$$M_{Fe}=55.895\,\mathrm{g/mol}$$

$$M_O=16.000\,\mathrm{g/mol}$$

$$M_c=159.69\,\mathrm{g/mol}$$

$$\sigma_{Fe}=138.860\,\mathrm{cm^2/g}$$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

begin by finding tabulated values of the cross-section for the elements Fe and O at $5\ \text{keV}$

$$\rho = 5.24\,\mathrm{g/cm^2}$$

$$M_{Fe} = 55.895\,\mathrm{g/mol}$$

$$M_O = 16.000\,\mathrm{g/mol}$$

$$M_c=159.69\,\mathrm{g/mol}$$

$$\sigma_{Fe}=138.860\,\mathrm{cm^2/g}$$

$$\sigma_O = 46.666 \, \mathrm{cm}^2/\mathrm{g}$$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

begin by finding tabulated values of the cross-section for the elements Fe and O at $5\ keV$

$$\rho = 5.24\,\mathrm{g/cm^2}$$

$$M_{Fe}=55.895\,\mathrm{g/mol}$$

$$M_O=16.000\,\mathrm{g/mol}$$

$$M_c=159.69\,\mathrm{g/mol}$$

$$\sigma_{Fe}=138.860\,\mathrm{cm^2/g}$$

$$\sigma_O = 46.666 \, \mathrm{cm}^2/\mathrm{g}$$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

begin by finding tabulated values of the cross-section for the elements Fe and O at $5\ \text{keV}$

$$\overline{M_c} \stackrel{\textstyle \sum_j}{} \overline{N_A}^{xjoj} - \overline{M_c} \stackrel{\textstyle \sum_j}{} \overline{M_j x_j o_j}$$
 $ho = 5.24 \, \mathrm{g/cm^2}$
 $M_{Fe} = 55.895 \, \mathrm{g/mol}$
 $M_O = 16.000 \, \mathrm{g/mol}$
 $M_c = 159.69 \, \mathrm{g/mol}$
 $\sigma_{Fe} = 138.860 \, \mathrm{cm^2/g}$
 $\sigma_O = 46.666 \, \mathrm{cm^2/g}$
 $A = \pi (0.25 \, \mathrm{cm})^2 = 0.1963 \, \mathrm{cm^2}$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

begin by finding tabulated values of the cross-section for the elements Fe and O at $5\ \text{keV}$

$$\rho = 5.24\,\mathrm{g/cm^2}$$

$$M_{Fe} = 55.895 \,\mathrm{g/mol}$$

$$M_O=16.000\,\mathrm{g/mol}$$

$$M_c=159.69\,\mathrm{g/mol}$$

$$\sigma_{Fe}=138.860\,\mathrm{cm^2/g}$$

$$\sigma_O = 46.666\,\mathrm{cm}^2/\mathrm{g}$$

$$A = \pi (0.25 \, \text{cm})^2 = 0.1963 \, \text{cm}^2$$

$$\mu_{m} = \frac{1}{159.69} \left[2 \cdot 55.895 \cdot 138.860 + 3 \cdot 16.000 \cdot 46.666 \right]$$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

begin by finding tabulated values of the cross-section for the elements Fe and O at $5\ \text{keV}$

$$\rho = 5.24\,\mathrm{g/cm^2}$$

$$M_{Fe} = 55.895\,\mathrm{g/mol}$$

$$M_O=16.000\,\mathrm{g/mol}$$

$$M_c=159.69\,\mathrm{g/mol}$$

$$\sigma_{Fe} = 138.860\,\mathrm{cm^2/g}$$

$$\sigma_O = 46.666\,\mathrm{cm}^2/\mathrm{g}$$

$$A = \pi (0.25 \, \text{cm})^2 = 0.1963 \, \text{cm}^2$$

$$\mu_{\it m} = {1 \over 159.69} \left[2 \cdot 55.895 \cdot 138.860 + 3 \cdot 16.000 \cdot 46.666 \right] = 111.23 \, {\rm cm}^2/{\rm g}$$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

begin by finding tabulated values of the cross-section for the elements Fe and O at $5\ \text{keV}$

assuming a 5 mm diameter pellet

$$\rho = 5.24 \,\mathrm{g/cm^2}$$

$$M_{Fe} = 55.895 \,\mathrm{g/mol}$$

$$M_O = 16.000\,\mathrm{g/mol}$$

$$M_c=159.69\,\mathrm{g/mol}$$

$$\sigma_{Fe} = 138.860\,\mathrm{cm^2/g}$$

$$\sigma_O = 46.666\,\mathrm{cm}^2/\mathrm{g}$$

$$A = \pi (0.25 \,\mathrm{cm})^2 = 0.1963 \,\mathrm{cm}^2$$

$$\mu_{m} = \frac{1}{159.69} \left[2 \cdot 55.895 \cdot 138.860 + 3 \cdot 16.000 \cdot 46.666 \right] = 111.23 \,\text{cm}^{2}/\text{g}$$

 $\mu_{\it m}/A = 566.7\,{\rm g}^{-1}$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_j = N_A \sigma_{aj}/M_j$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

begin by finding tabulated values of the cross-section for the elements Fe and O at 5 keV

$$\rho = 5.24\,\mathrm{g/cm^2}$$

$$M_{Fe} = 55.895 \,\mathrm{g/mol}$$

$$M_O = 16.000\,\mathrm{g/mol}$$

$$M_c=159.69\,\mathrm{g/mol}$$

$$\sigma_{Fe} = 138.860\,\mathrm{cm^2/g}$$

$$\sigma_O = 46.666\,\mathrm{cm}^2/\mathrm{g}$$

$$A = \pi (0.25 \, \text{cm})^2 = 0.1963 \, \text{cm}^2$$

$$\mu_{m} = \frac{1}{159.69} \left[2 \cdot 55.895 \cdot 138.860 + 3 \cdot 16.000 \cdot 46.666 \right] = 111.23 \,\text{cm}^{2}/\text{g}$$

$$\mu_{\rm m}/A = 566.7 \,{\rm g}^{-1}$$
 $\mu = \mu_{\rm m} \rho = 582.9 \,{\rm cm}^{-1}$

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_i = N_A \sigma_{ai}/M_i$ so we have

$$I = I_0 e^{-(\mu_m/A)m}, \ \mu_m = \frac{N_A}{M_c} \sum_j x_j \sigma_{aj} = \frac{N_A}{M_c} \sum_j \frac{M_j}{N_A} x_j \sigma_j = \frac{1}{M_c} \sum_j M_j x_j \sigma_j$$

the molecular mass and density of Fe_2O_3 are

begin by finding tabulated values of the cross-section for the elements Fe and O at 5 keV

$$\rho = 5.24\,\mathrm{g/cm^2}$$

$$M_{Fe} = 55.895 \,\mathrm{g/mol}$$

$$M_O=16.000\,\mathrm{g/mol}$$

$$M_c=159.69\,\mathrm{g/mol}$$

$$\sigma_{Fe}=138.860\,\mathrm{cm^2/g}$$

$$\sigma_O = 46.666\,\mathrm{cm}^2/\mathrm{g}$$

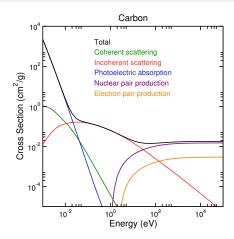
$$A = \pi (0.25 \, \text{cm})^2 = 0.1963 \, \text{cm}^2$$

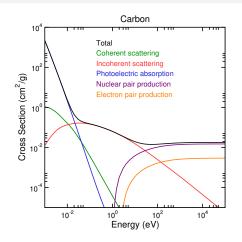
$$\mu_{m} = \frac{1}{159.69} \left[2 \cdot 55.895 \cdot 138.860 + 3 \cdot 16.000 \cdot 46.666 \right] = 111.23 \,\text{cm}^{2}/\text{g}$$

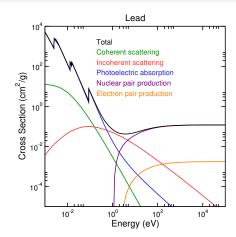
$$\mu_{\rm m}/A = 566.7 \,{\rm g}^{-1}$$

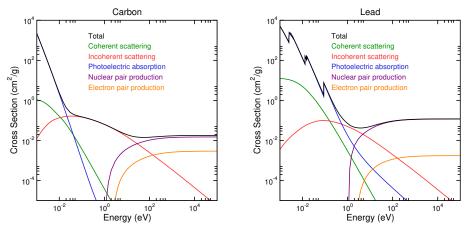
$$\mu_{\rm m}/A = 566.7 \,{\rm g}^{-1}$$
 $\mu = \mu_{\rm m} \rho = 582.9 \,{\rm cm}^{-1}$ $1/\mu = 17.2 \,\mu{\rm m}$

$$1/\mu = 17.2 \, \mu \text{m}$$

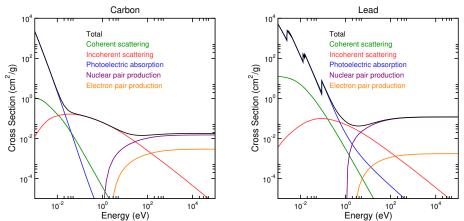






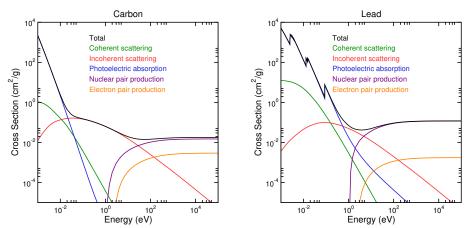


Photoelectric absorption dominates at low energies

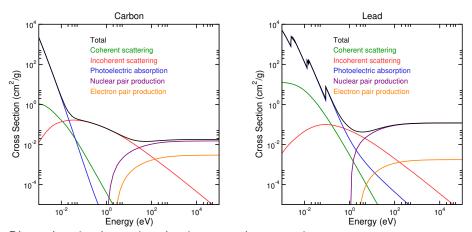


Photoelectric absorption dominates at low energies

Thomson scattering (coherent) drops rapidly with energy



Photoelectric absorption dominates at low energies Thomson scattering (coherent) drops rapidly with energy Compton scattering (incoherent) dominates at medium energies



Photoelectric absorption dominates at low energies Thomson scattering (coherent) drops rapidly with energy Compton scattering (incoherent) dominates at medium energies Pair production dominates at high energies