## Today's outline - January 16, 2020

## Today's outline - January 16, 2020

- Crystal lattice types


## Today's outline - January 16, 2020

- Crystal lattice types
- The reciprocal lattice


## Today's outline - January 16, 2020

- Crystal lattice types
- The reciprocal lattice
- Compton (inelastic) scattering


## Today's outline - January 16, 2020

- Crystal lattice types
- The reciprocal lattice
- Compton (inelastic) scattering
- X-ray absorption


## Scattering from a crystal (review)

Recall that for a crystal lattice which is a periodic array of molecules

## Scattering from a crystal (review)

Recall that for a crystal lattice which is a periodic array of molecules


$$
F^{\text {crystal }}(\mathbf{Q})=F^{\text {molecule }} F^{\text {lattice }}
$$

## Scattering from a crystal (review)

Recall that for a crystal lattice which is a periodic array of molecules


$$
\begin{array}{r}
F^{\text {crystal }}(\mathbf{Q})=F^{\text {molecule }} F^{\text {lattice }} \\
F^{\text {crystal }}(\mathbf{Q})=\sum_{j} f_{j}(\mathbf{Q}) e^{i \mathbf{Q} \cdot \mathbf{r}_{j}} \sum_{n} e^{i \mathbf{Q} \cdot \mathbf{R}_{n}}
\end{array}
$$

## Scattering from a crystal (review)

Recall that for a crystal lattice which is a periodic array of molecules


$$
\begin{array}{r}
F^{\text {crystal }}(\mathbf{Q})=F^{\text {molecule }} F^{\text {lattice }} \\
F^{\text {crystal }}(\mathbf{Q})=\sum_{j} f_{j}(\mathbf{Q}) e^{i \mathbf{Q} \cdot \mathbf{r}_{j}} \sum_{n} e^{i \mathbf{Q} \cdot \mathbf{R}_{n}}
\end{array}
$$

The lattice term, $\sum e^{i \mathbf{Q} \cdot \mathbf{R}_{n}}$, is a sum over a large number

## Scattering from a crystal (review)

Recall that for a crystal lattice which is a periodic array of molecules


$$
\begin{array}{r}
F^{\text {crystal }}(\mathbf{Q})=F^{\text {molecule }} F^{\text {lattice }} \\
F^{\text {crystal }}(\mathbf{Q})=\sum_{j} f_{j}(\mathbf{Q}) e^{i \mathbf{Q} \cdot \mathbf{r}_{j}} \sum_{n} e^{i \mathbf{Q} \cdot \mathbf{R}_{n}}
\end{array}
$$

The lattice term, $\sum e^{i \mathbf{Q} \cdot \mathbf{R}_{n}}$, is a sum over a large number so it is always small unless $\mathbf{Q} \cdot \mathbf{R}_{n}=2 \pi m$ where $\mathbf{R}_{n}=n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}$ is a real space lattice vector and $m$ is an integer.

## Scattering from a crystal (review)

Recall that for a crystal lattice which is a periodic array of molecules


$$
\begin{array}{r}
F^{\text {crystal }}(\mathbf{Q})=F^{\text {molecule }} F^{\text {lattice }} \\
F^{\text {crystal }}(\mathbf{Q})=\sum_{j} f_{j}(\mathbf{Q}) e^{i \mathbf{Q} \cdot \mathbf{r}_{j}} \sum_{n} e^{i \mathbf{Q} \cdot \mathbf{R}_{n}}
\end{array}
$$

The lattice term, $\sum e^{i \mathbf{Q} \cdot \mathbf{R}_{n}}$, is a sum over a large number so it is always small unless $\mathbf{Q} \cdot \mathbf{R}_{n}=2 \pi m$ where $\mathbf{R}_{n}=n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}$ is a real space lattice vector and $m$ is an integer.

This condition is fulfilled only when $\mathbf{Q}$ is a reciprocal lattice vector.

## Crystal lattices

There are 7 possible real space lattices: triclinic,


## Crystal lattices

There are 7 possible real space lattices: triclinic, monoclinic,


## Crystal lattices

There are 7 possible real space lattices: triclinic, monoclinic, orthorhombic,


## Crystal lattices

There are 7 possible real space lattices: triclinic, monoclinic, orthorhombic, tetragonal,


## Crystal lattices

There are 7 possible real space lattices: triclinic, monoclinic, orthorhombic, tetragonal, hexagonal,


## Crystal lattices

There are 7 possible real space lattices: triclinic, monoclinic, orthorhombic, tetragonal, hexagonal, rhombohedral,


## Crystal lattices

There are 7 possible real space lattices: triclinic, monoclinic, orthorhombic, tetragonal, hexagonal, rhombohedral, cubic


## Lattice properties

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).


## Lattice properties

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).


$$
\mathbf{a}_{1}=a \hat{\mathbf{x}}, \quad \mathbf{a}_{2}=b \hat{\mathbf{y}}, \quad \mathbf{a}_{3}=c \hat{\mathbf{z}}
$$

## Lattice properties

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).


$$
\begin{aligned}
\mathbf{a}_{1}=a \hat{\mathbf{x}}, \quad \mathbf{a}_{2} & =b \hat{\mathbf{y}}, \quad \mathbf{a}_{3}=c \hat{\mathbf{z}} \\
\mathbf{a}_{1} \times \mathbf{a}_{2} & =a b \hat{\mathbf{z}}
\end{aligned}
$$

## Lattice properties

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).


$$
\begin{aligned}
\mathbf{a}_{1}=a \hat{\mathbf{x}}, \quad \mathbf{a}_{2} & =b \hat{\mathbf{y}}, \quad \mathbf{a}_{3}=c \hat{\mathbf{z}} \\
\mathbf{a}_{1} \times \mathbf{a}_{2} & =a b \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b \hat{\mathbf{z}} \cdot c \hat{\mathbf{z}}
\end{aligned}
$$

## Lattice properties

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).


$$
\begin{aligned}
\mathbf{a}_{1}=a \hat{\mathbf{x}}, \quad \mathbf{a}_{2} & =b \hat{\mathbf{y}}, \quad \mathbf{a}_{3}=c \hat{\mathbf{z}} \\
\mathbf{a}_{1} \times \mathbf{a}_{2} & =a b \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b \hat{\mathbf{z}} \cdot c \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b c=V
\end{aligned}
$$

## Lattice properties

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).


$$
\begin{aligned}
\mathbf{a}_{1}=a \hat{\mathbf{x}}, \quad \mathbf{a}_{2} & =b \hat{\mathbf{y}}, \quad \mathbf{a}_{3}=c \hat{\mathbf{z}} \\
\mathbf{a}_{1} \times \mathbf{a}_{2} & =a b \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b \hat{\mathbf{z}} \cdot c \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b c=V
\end{aligned}
$$

A simple way of calculating the volume of the unit cell!

## Lattice properties

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).


$$
\begin{aligned}
\mathbf{a}_{1}=a \hat{\mathbf{x}}, \quad \mathbf{a}_{2} & =b \hat{\mathbf{y}}, \quad \mathbf{a}_{3}=c \hat{\mathbf{z}} \\
\mathbf{a}_{1} \times \mathbf{a}_{2} & =a b \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b \hat{\mathbf{z}} \cdot c \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b c=V
\end{aligned}
$$

A simple way of calculating the volume of the unit cell!
This unit cell is repeated infintely in 3-dimensions and thus, the location of each lattice point can be calculated relative to any arbitrary lattice point designated as the origin.

## Lattice properties

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).


$$
\begin{aligned}
\mathbf{a}_{1}=a \hat{\mathbf{x}}, \quad \mathbf{a}_{2} & =b \hat{\mathbf{y}}, \quad \mathbf{a}_{3}=c \hat{\mathbf{z}} \\
\mathbf{a}_{1} \times \mathbf{a}_{2} & =a b \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b \hat{\mathbf{z}} \cdot c \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b c=V
\end{aligned}
$$

A simple way of calculating the volume of the unit cell!
This unit cell is repeated infintely in 3-dimensions and thus, the location of each lattice point can be calculated relative to any arbitrary lattice point designated as the origin.

Each lattice point is at the end of a lattice vector, $\mathbf{R}_{n}$

## Lattice properties

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).


$$
\begin{aligned}
\mathbf{a}_{1}=a \hat{\mathbf{x}}, \quad \mathbf{a}_{2} & =b \hat{\mathbf{y}}, \quad \mathbf{a}_{3}=c \hat{\mathbf{z}} \\
\mathbf{a}_{1} \times \mathbf{a}_{2} & =a b \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b \hat{\mathbf{z}} \cdot c \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b c=V
\end{aligned}
$$

A simple way of calculating the volume of the unit cell!
This unit cell is repeated infintely in 3-dimensions and thus, the location of each lattice point can be calculated relative to any arbitrary lattice point designated as the origin.

Each lattice point is at the end of a lattice vector, $\mathbf{R}_{n}$

$$
\mathbf{R}_{n}=n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}
$$

## Lattice properties

Consider the orthorhombic lattice for simplicity (the others give exactly the same result).


$$
\begin{aligned}
\mathbf{a}_{1}=a \hat{\mathbf{x}}, \quad \mathbf{a}_{2} & =b \hat{\mathbf{y}}, \quad \mathbf{a}_{3}=c \hat{\mathbf{z}} \\
\mathbf{a}_{1} \times \mathbf{a}_{2} & =a b \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b \hat{\mathbf{z}} \cdot c \hat{\mathbf{z}} \\
\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right) \cdot \mathbf{a}_{3} & =a b c=V
\end{aligned}
$$

A simple way of calculating the volume of the unit cell!
This unit cell is repeated infintely in 3-dimensions and thus, the location of each lattice point can be calculated relative to any arbitrary lattice point designated as the origin.

Each lattice point is at the end of a lattice vector, $\mathbf{R}_{n}$ and a crystal is made by putting a molecule at each lattice point.

$$
\mathbf{R}_{n}=n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}
$$

## Reciprocal lattice

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

## Reciprocal lattice

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

## Reciprocal lattice

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

$$
\mathbf{a}_{1}^{*}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot\left(\mathbf{a}_{2} \times \mathbf{a}_{3}\right)}
$$

## Reciprocal lattice

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

$$
\begin{aligned}
& \mathbf{a}_{1}^{*}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot\left(\mathbf{a}_{2} \times \mathbf{a}_{3}\right)} \\
& \mathbf{a}_{2}^{*}=2 \pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{\mathbf{a}_{2} \cdot\left(\mathbf{a}_{3} \times \mathbf{a}_{1}\right)}
\end{aligned}
$$

## Reciprocal lattice

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

$$
\begin{aligned}
& \mathbf{a}_{1}^{*}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot\left(\mathbf{a}_{2} \times \mathbf{a}_{3}\right)} \\
& \mathbf{a}_{2}^{*}=2 \pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{\mathbf{a}_{2} \cdot\left(\mathbf{a}_{3} \times \mathbf{a}_{1}\right)} \\
& \mathbf{a}_{3}^{*}=2 \pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{\mathbf{a}_{3} \cdot\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right)}
\end{aligned}
$$

## Reciprocal lattice

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

$$
\begin{aligned}
& \mathbf{a}_{1}^{*}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot\left(\mathbf{a}_{2} \times \mathbf{a}_{3}\right)}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} \\
& \mathbf{a}_{2}^{*}=2 \pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{\mathbf{a}_{2} \cdot\left(\mathbf{a}_{3} \times \mathbf{a}_{1}\right)}=2 \pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} \\
& \mathbf{a}_{3}^{*}=2 \pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{\mathbf{a}_{3} \cdot\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right)}=2 \pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}
\end{aligned}
$$

## Reciprocal lattice

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

$$
\begin{aligned}
& \mathbf{a}_{1}^{*}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot\left(\mathbf{a}_{2} \times \mathbf{a}_{3}\right)}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} \\
& \mathbf{a}_{2}^{*}=2 \pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{\mathbf{a}_{2} \cdot\left(\mathbf{a}_{3} \times \mathbf{a}_{1}\right)}=2 \pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} \\
& \mathbf{a}_{3}^{*}=2 \pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{\mathbf{a}_{3} \cdot\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right)}=2 \pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}
\end{aligned}
$$

In analogy to $\mathbf{R}_{n}$, we can construct an arbitrary reciprocal space lattice vector $\mathbf{G}_{h k l}$

## Reciprocal lattice

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

$$
\begin{aligned}
& \mathbf{a}_{1}^{*}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot\left(\mathbf{a}_{2} \times \mathbf{a}_{3}\right)}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} \\
& \mathbf{a}_{2}^{*}=2 \pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{\mathbf{a}_{2} \cdot\left(\mathbf{a}_{3} \times \mathbf{a}_{1}\right)}=2 \pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} \\
& \mathbf{a}_{3}^{*}=2 \pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{\mathbf{a}_{3} \cdot\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right)}=2 \pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}
\end{aligned}
$$

In analogy to $\mathbf{R}_{n}$, we can construct an arbitrary reciprocal space lattice vector $\mathbf{G}_{h k l}$

$$
\mathbf{G}_{h k l}=h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}
$$

## Reciprocal lattice

For any lattice in real space, it is useful to construct what is called a reciprocal space lattice.

Define the reciprocal lattice vectors in terms of the real space unit vectors

$$
\begin{aligned}
& \mathbf{a}_{1}^{*}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{\mathbf{a}_{1} \cdot\left(\mathbf{a}_{2} \times \mathbf{a}_{3}\right)}=2 \pi \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V} \\
& \mathbf{a}_{2}^{*}=2 \pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{\mathbf{a}_{2} \cdot\left(\mathbf{a}_{3} \times \mathbf{a}_{1}\right)}=2 \pi \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V} \\
& \mathbf{a}_{3}^{*}=2 \pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{\mathbf{a}_{3} \cdot\left(\mathbf{a}_{1} \times \mathbf{a}_{2}\right)}=2 \pi \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}
\end{aligned}
$$

In analogy to $\mathbf{R}_{n}$, we can construct an arbitrary reciprocal space lattice vector $\mathbf{G}_{h k l}$

$$
\mathbf{G}_{h k l}=h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}
$$

where $h, k$, and $I$ are integers called Miller indices

## Laue condition

Because of the construction of the reciprocal lattice

## Laue condition

Because of the construction of the reciprocal lattice
$\mathbf{G}_{n k l} \cdot \mathbf{R}_{n}$

## Laue condition

Because of the construction of the reciprocal lattice

$$
\mathbf{G}_{h k l} \cdot \mathbf{R}_{n}=\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot\left(h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}\right)
$$

## Laue condition

Because of the construction of the reciprocal lattice

$$
\begin{aligned}
\mathbf{G}_{h k l} \cdot \mathbf{R}_{n} & =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot\left(h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}\right) \\
& =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot 2 \pi\left(h \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}+k \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}+I \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right)
\end{aligned}
$$

## Laue condition

Because of the construction of the reciprocal lattice

$$
\begin{aligned}
\mathbf{G}_{h k l} \cdot \mathbf{R}_{n} & =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot\left(h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}\right) \\
& =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot 2 \pi\left(h \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}+k \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}+I \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right) \\
& =2 \pi\left(h n_{1}+k n_{2}+l n_{3}\right)=2 \pi m
\end{aligned}
$$

## Laue condition

Because of the construction of the reciprocal lattice

$$
\begin{aligned}
\mathbf{G}_{h k l} \cdot \mathbf{R}_{n} & =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot\left(h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}\right) \\
& =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot 2 \pi\left(h \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}+k \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}+I \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right) \\
& =2 \pi\left(h n_{1}+k n_{2}+l n_{3}\right)=2 \pi m
\end{aligned}
$$

and therefore, the crystal scattering factor

## Laue condition

Because of the construction of the reciprocal lattice

$$
\begin{aligned}
\mathbf{G}_{h k l} \cdot \mathbf{R}_{n} & =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot\left(h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}\right) \\
& =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot 2 \pi\left(h \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}+k \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}+I \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right) \\
& =2 \pi\left(h n_{1}+k n_{2}+l n_{3}\right)=2 \pi m
\end{aligned}
$$

and therefore, the crystal scattering factor

$$
\sum e^{i \mathbf{Q} \cdot \mathbf{R}_{n}}
$$

## Laue condition

Because of the construction of the reciprocal lattice

$$
\begin{aligned}
\mathbf{G}_{h k l} \cdot \mathbf{R}_{n} & =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot\left(h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}\right) \\
& =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot 2 \pi\left(h \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}+k \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}+I \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right) \\
& =2 \pi\left(h n_{1}+k n_{2}+l n_{3}\right)=2 \pi m
\end{aligned}
$$

and therefore, the crystal scattering factor is non-zero only when

$$
\sum e^{i \mathbf{Q} \cdot \mathbf{R}_{n}} \neq 0
$$

## Laue condition

Because of the construction of the reciprocal lattice

$$
\begin{aligned}
\mathbf{G}_{h k l} \cdot \mathbf{R}_{n} & =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot\left(h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}\right) \\
& =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot 2 \pi\left(h \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}+k \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}+I \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right) \\
& =2 \pi\left(h n_{1}+k n_{2}+l n_{3}\right)=2 \pi m
\end{aligned}
$$

and therefore, the crystal scattering factor is non-zero only when

$$
\sum e^{i \mathbf{Q} \cdot \mathbf{R}_{n}} \neq 0 \quad \mathbf{Q}=\mathbf{G}_{h k l}
$$

## Laue condition

Because of the construction of the reciprocal lattice

$$
\begin{aligned}
\mathbf{G}_{h k l} \cdot \mathbf{R}_{n} & =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot\left(h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}\right) \\
& =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot 2 \pi\left(h \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}+k \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}+I \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right) \\
& =2 \pi\left(h n_{1}+k n_{2}+l n_{3}\right)=2 \pi m
\end{aligned}
$$

and therefore, the crystal scattering factor is non-zero only when

$$
\sum e^{i \mathbf{Q} \cdot \mathbf{R}_{n}} \neq 0 \quad \mathbf{Q}=\mathbf{G}_{h k l}
$$

so a significant number of molecules scatter in phase with each other

## Laue condition

Because of the construction of the reciprocal lattice

$$
\begin{aligned}
\mathbf{G}_{h k l} \cdot \mathbf{R}_{n} & =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot\left(h \mathbf{a}_{1}^{*}+k \mathbf{a}_{2}^{*}+l \mathbf{a}_{3}^{*}\right) \\
& =\left(n_{1} \mathbf{a}_{1}+n_{2} \mathbf{a}_{2}+n_{3} \mathbf{a}_{3}\right) \cdot 2 \pi\left(h \frac{\mathbf{a}_{2} \times \mathbf{a}_{3}}{V}+k \frac{\mathbf{a}_{3} \times \mathbf{a}_{1}}{V}+I \frac{\mathbf{a}_{1} \times \mathbf{a}_{2}}{V}\right) \\
& =2 \pi\left(h n_{1}+k n_{2}+l n_{3}\right)=2 \pi m
\end{aligned}
$$

and therefore, the crystal scattering factor is non-zero only when

$$
\sum e^{i \mathbf{Q} \cdot \mathbf{R}_{n}} \neq 0 \quad \mathbf{Q}=\mathbf{G}_{h k l}
$$

so a significant number of molecules scatter in phase with each other
As we shall see later, this Laue condition, is equivalent to the more typically used Bragg condition for diffraction: $2 d \sin \theta=n \lambda$

## Multiple slit interference

A crystal is, therefore, a diffraction grating with $\sim 10^{20}$ slits!

## Multiple slit interference

A crystal is, therefore, a diffraction grating with $\sim 10^{20}$ slits!
When $\mathbf{Q}$ is a reciprocal lattice vector, a very strong, narrow diffraction peak is seen at the detector.

## Multiple slit interference

A crystal is, therefore, a diffraction grating with $\sim 10^{20}$ slits!
When $\mathbf{Q}$ is a reciprocal lattice vector, a very strong, narrow diffraction peak is seen at the detector.


## Multiple slit interference

A crystal is, therefore, a diffraction grating with $\sim 10^{20}$ slits!
When $\mathbf{Q}$ is a reciprocal lattice vector, a very strong, narrow diffraction peak is seen at the detector.



## Compton scattering

A photon-electron collision

## Compton scattering

A photon-electron collision


## Compton scattering

A photon-electron collision


$$
\mathbf{p}=\hbar \mathbf{k}=2 \pi \hbar / \lambda
$$

## Compton scattering

A photon-electron collision


$$
\begin{array}{r}
\mathbf{p}=\hbar \mathbf{k}=2 \pi \hbar / \lambda \\
\mathbf{p}^{\prime}=\hbar \mathbf{k}^{\prime}=2 \pi \hbar / \lambda^{\prime}
\end{array}
$$

## Compton scattering

A photon-electron collision


$$
\begin{array}{r}
\mathbf{p}=\hbar \mathbf{k}=2 \pi \hbar / \lambda \\
\mathbf{p}^{\prime}=\hbar \mathbf{k}^{\prime}=2 \pi \hbar / \lambda^{\prime} \\
|\mathbf{k}| \neq\left|\mathbf{k}^{\prime}\right|
\end{array}
$$

## Compton scattering

A photon-electron collision


$$
\begin{array}{r}
\mathbf{p}=\hbar \mathbf{k}=2 \pi \hbar / \lambda \\
\mathbf{p}^{\prime}=\hbar \mathbf{k}^{\prime}=2 \pi \hbar / \lambda^{\prime} \\
|\mathbf{k}| \neq\left|\mathbf{k}^{\prime}\right|
\end{array}
$$

Treat the electron relativistically and conserve energy and momentum

## Compton scattering

A photon-electron collision


$$
\begin{array}{r}
\mathbf{p}=\hbar \mathbf{k}=2 \pi \hbar / \lambda \\
\mathbf{p}^{\prime}=\hbar \mathbf{k}^{\prime}=2 \pi \hbar / \lambda^{\prime} \\
|\mathbf{k}| \neq\left|\mathbf{k}^{\prime}\right|
\end{array}
$$

Treat the electron relativistically and conserve energy and momentum

$$
m c^{2}+\frac{h c}{\lambda}=\frac{h c}{\lambda^{\prime}}+\gamma m c^{2} \quad \text { (energy) }
$$

## Compton scattering

A photon-electron collision


$$
\begin{array}{r}
\mathbf{p}=\hbar \mathbf{k}=2 \pi \hbar / \lambda \\
\mathbf{p}^{\prime}=\hbar \mathbf{k}^{\prime}=2 \pi \hbar / \lambda^{\prime} \\
|\mathbf{k}| \neq\left|\mathbf{k}^{\prime}\right|
\end{array}
$$

Treat the electron relativistically and conserve energy and momentum

$$
\begin{array}{ll}
m c^{2}+\frac{h c}{\lambda}=\frac{h c}{\lambda^{\prime}}+\gamma m c^{2} & \text { (energy) } \\
\frac{h}{\lambda}=\frac{h}{\lambda^{\prime}} \cos \phi+\gamma m v \cos \theta & (x \text {-axis })
\end{array}
$$

## Compton scattering

A photon-electron collision


$$
\begin{array}{r}
\mathbf{p}=\hbar \mathbf{k}=2 \pi \hbar / \lambda \\
\mathbf{p}^{\prime}=\hbar \mathbf{k}^{\prime}=2 \pi \hbar / \lambda^{\prime} \\
|\mathbf{k}| \neq\left|\mathbf{k}^{\prime}\right|
\end{array}
$$

Treat the electron relativistically and conserve energy and momentum

$$
\begin{array}{cc}
m c^{2}+\frac{h c}{\lambda}=\frac{h c}{\lambda^{\prime}}+\gamma m c^{2} & (\text { energy }) \\
\frac{h}{\lambda}=\frac{h}{\lambda^{\prime}} \cos \phi+\gamma m v \cos \theta & (x \text {-axis }) \\
0=\frac{h}{\lambda^{\prime}} \sin \phi+\gamma m v \sin \theta & (y \text {-axis })
\end{array}
$$

## Compton scattering derivation

squaring the momentum equations

## Compton scattering derivation

squaring the momentum $\quad\left(\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}} \cos \phi\right)^{2}=\gamma^{2} m^{2} v^{2} \cos ^{2} \theta$
equations

## Compton scattering derivation

squaring the momentum

$$
\begin{aligned}
\left(\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}} \cos \phi\right)^{2} & =\gamma^{2} m^{2} v^{2} \cos ^{2} \theta \\
\left(-\frac{h}{\lambda^{\prime}} \sin \phi\right)^{2} & =\gamma^{2} m^{2} v^{2} \sin ^{2} \theta
\end{aligned}
$$

## Compton scattering derivation

squaring the momentum

$$
\begin{aligned}
\left(\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}} \cos \phi\right)^{2} & =\gamma^{2} m^{2} v^{2} \cos ^{2} \theta \\
\left(-\frac{h}{\lambda^{\prime}} \sin \phi\right)^{2} & =\gamma^{2} m^{2} v^{2} \sin ^{2} \theta
\end{aligned}
$$

now add them together,

$$
\gamma^{2} m^{2} v^{2}\left(\sin ^{2} \theta+\cos ^{2} \theta\right)=\left(\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}} \cos \phi\right)^{2}+\left(-\frac{h}{\lambda^{\prime}} \sin \phi\right)^{2}
$$

## Compton scattering derivation

squaring the momentum equations

$$
\begin{aligned}
\left(\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}} \cos \phi\right)^{2} & =\gamma^{2} m^{2} v^{2} \cos ^{2} \theta \\
\left(-\frac{h}{\lambda^{\prime}} \sin \phi\right)^{2} & =\gamma^{2} m^{2} v^{2} \sin ^{2} \theta
\end{aligned}
$$

now add them together, substitute $\sin ^{2} \theta+\cos ^{2} \theta=1$, expand the squares,

$$
\begin{aligned}
\gamma^{2} m^{2} v^{2}\left(\sin ^{2} \theta+\cos ^{2} \theta\right) & =\left(\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}} \cos \phi\right)^{2}+\left(-\frac{h}{\lambda^{\prime}} \sin \phi\right)^{2} \\
\gamma^{2} m^{2} v^{2} & =\frac{h^{2}}{\lambda^{2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi+\frac{h^{2}}{\lambda^{\prime 2}} \sin ^{2} \phi+\frac{h^{2}}{\lambda^{\prime 2}} \cos ^{2} \phi
\end{aligned}
$$

## Compton scattering derivation

squaring the momentum equations

$$
\begin{aligned}
\left(\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}} \cos \phi\right)^{2} & =\gamma^{2} m^{2} v^{2} \cos ^{2} \theta \\
\left(-\frac{h}{\lambda^{\prime}} \sin \phi\right)^{2} & =\gamma^{2} m^{2} v^{2} \sin ^{2} \theta
\end{aligned}
$$

now add them together, substitute $\sin ^{2} \theta+\cos ^{2} \theta=1$, expand the squares, and $\sin ^{2} \phi+\cos ^{2} \phi=1$, then rearrange

$$
\begin{aligned}
\gamma^{2} m^{2} v^{2}\left(\sin ^{2} \theta+\cos ^{2} \theta\right) & =\left(\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}} \cos \phi\right)^{2}+\left(-\frac{h}{\lambda^{\prime}} \sin \phi\right)^{2} \\
\gamma^{2} m^{2} v^{2} & =\frac{h^{2}}{\lambda^{2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi+\frac{h^{2}}{\lambda^{\prime 2}} \sin ^{2} \phi+\frac{h^{2}}{\lambda^{\prime 2}} \cos ^{2} \phi \\
\frac{m^{2} v^{2}}{1-\beta^{2}} & =\frac{h^{2}}{\lambda^{2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi+\frac{h^{2}}{\lambda^{\prime 2}}
\end{aligned}
$$

## Compton scattering derivation

squaring the momentum equations

$$
\begin{aligned}
\left(\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}} \cos \phi\right)^{2} & =\gamma^{2} m^{2} v^{2} \cos ^{2} \theta \\
\left(-\frac{h}{\lambda^{\prime}} \sin \phi\right)^{2} & =\gamma^{2} m^{2} v^{2} \sin ^{2} \theta
\end{aligned}
$$

now add them together, substitute $\sin ^{2} \theta+\cos ^{2} \theta=1$, expand the squares, and $\sin ^{2} \phi+\cos ^{2} \phi=1$, then rearrange and substitute $v=\beta c$

$$
\begin{aligned}
\gamma^{2} m^{2} v^{2}\left(\sin ^{2} \theta+\cos ^{2} \theta\right) & =\left(\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}} \cos \phi\right)^{2}+\left(-\frac{h}{\lambda^{\prime}} \sin \phi\right)^{2} \\
\gamma^{2} m^{2} v^{2} & =\frac{h^{2}}{\lambda^{2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi+\frac{h^{2}}{\lambda^{\prime 2}} \sin ^{2} \phi+\frac{h^{2}}{\lambda^{\prime 2}} \cos ^{2} \phi \\
\frac{m^{2} c^{2} \beta^{2}}{1-\beta^{2}}=\frac{m^{2} v^{2}}{1-\beta^{2}} & =\frac{h^{2}}{\lambda^{2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi+\frac{h^{2}}{\lambda^{\prime 2}}
\end{aligned}
$$

## Compton scattering derivation (cont.)

Now take the energy equation and square it,

$$
\left(m c^{2}+\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)^{2}=\gamma^{2} m^{2} c^{4}=\frac{m^{2} c^{4}}{1-\beta^{2}}
$$

## Compton scattering derivation (cont.)

Now take the energy equation and square it, then solve it for $\beta^{2}$

$$
\begin{gathered}
\left(m c^{2}+\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)^{2}=\gamma^{2} m^{2} c^{4}=\frac{m^{2} c^{4}}{1-\beta^{2}} \\
\beta^{2}=1-\frac{m^{2} c^{4}}{\left(m c^{2}+\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)^{2}}
\end{gathered}
$$

## Compton scattering derivation (cont.)

Now take the energy equation and square it, then solve it for $\beta^{2}$ which is substituted into the equation from the momentum.

$$
\begin{gathered}
\left(m c^{2}+\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)^{2}=\gamma^{2} m^{2} c^{4}=\frac{m^{2} c^{4}}{1-\beta^{2}} \\
\beta^{2}=1-\frac{m^{2} c^{4}}{\left(m c^{2}+\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)^{2}} \\
\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi=\frac{m^{2} c^{2} \beta^{2}}{1-\beta^{2}}
\end{gathered}
$$

## Compton scattering derivation (cont.)

Now take the energy equation and square it, then solve it for $\beta^{2}$ which is substituted into the equation from the momentum.

$$
\begin{aligned}
&\left(m c^{2}+\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)^{2}=\gamma^{2} m^{2} c^{4}=\frac{m^{2} c^{4}}{1-\beta^{2}} \\
& \beta^{2}=1-\frac{m^{2} c^{4}}{\left(m c^{2}+\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)^{2}} \\
& \begin{aligned}
\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi & =\frac{m^{2} c^{2} \beta^{2}}{1-\beta^{2}} \\
& =\frac{1}{c^{2}}\left(m c^{2}+\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)^{2}-m^{2} c^{2}
\end{aligned}
\end{aligned}
$$

## Compton scattering derivation (cont.)

$$
\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi=\left(m c+\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}}\right)^{2}-m^{2} c^{2}
$$

## Compton scattering derivation (cont.)

After expansion,

$$
\begin{aligned}
\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi & =\left(m c+\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}}\right)^{2}-m^{2} c^{2} \\
& =m^{2} c^{2}+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 m c h}{\lambda}-\frac{2 m c h}{\lambda^{\prime}}+\frac{2 h^{2}}{\lambda \lambda^{\prime}}-m^{2} c^{2}
\end{aligned}
$$

## Compton scattering derivation (cont.)

After expansion, cancellation,

$$
\begin{aligned}
\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi & =\left(m c+\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}}\right)^{2}-m^{2} c^{2} \\
& =m^{2} c^{2}+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 m c h}{\lambda}-\frac{2 m c h}{\lambda^{\prime}}+\frac{2 h^{2}}{\lambda \lambda^{\prime}}-m^{2} c^{2} \\
& =2 m\left(\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}}
\end{aligned}
$$

## Compton scattering derivation (cont.)

After expansion, cancellation,

$$
\begin{aligned}
\frac{h^{2}}{\lambda^{2}}+\frac{h^{2} /}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi & =\left(m c+\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}}\right)^{2}-m^{2} c^{2} \\
& =m^{2} c^{2}+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 m c h}{\lambda}-\frac{2 m c h}{\lambda^{\prime}}+\frac{2 h^{2}}{\lambda \lambda^{\prime}}-m^{2} c^{2} \\
& =2 m\left(\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)+\frac{h^{2} /}{\lambda^{2}}+\frac{h^{2} /}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}}
\end{aligned}
$$

## Compton scattering derivation (cont.)

After expansion, cancellation, and rearrangement, we obtain

$$
\begin{aligned}
\frac{h^{2}}{\lambda^{2}}+\frac{h^{2} /}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi & =\left(m c+\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}}\right)^{2}-m^{2} c^{2} \\
& =m^{2} c^{2}+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 m c h}{\lambda}-\frac{2 m c h}{\lambda^{\prime}}+\frac{2 h^{2}}{\lambda \lambda^{\prime}}-m^{2} c^{2} \\
& =2 m\left(\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)+\frac{h^{2} /}{\lambda^{2}}+\frac{h^{2} /}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \\
\frac{2 h^{2}}{\lambda \lambda^{\prime}}(1-\cos \phi) & =2 m\left(\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)
\end{aligned}
$$

## Compton scattering derivation (cont.)

After expansion, cancellation,

$$
\begin{aligned}
\frac{h^{2}}{\lambda^{2}}+\frac{h^{2} /}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi & =\left(m c+\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}}\right)^{2}-m^{2} c^{2} \\
& =m^{2} c^{2}+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 m c h}{\lambda}-\frac{2 m c h}{\lambda^{\prime}}+\frac{2 h^{2}}{\lambda \lambda^{\prime}}-m^{2} c^{2} \\
& =2 m\left(\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2} /}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \\
\frac{2 h^{2}}{\lambda \lambda^{\prime}}(1-\cos \phi) & =2 m\left(\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)=2 m h c\left(\frac{\lambda^{\prime}-\lambda}{\lambda \lambda^{\prime}}\right)
\end{aligned}
$$

## Compton scattering derivation (cont.)

After expansion, cancellation,

$$
\begin{aligned}
\frac{h^{2}}{\lambda^{2}}+\frac{h^{2} /}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi & =\left(m c+\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}}\right)^{2}-m^{2} c^{2} \\
& =m^{2} c^{2}+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 m c h}{\lambda}-\frac{2 m c h}{\lambda^{\prime}}+\frac{2 h^{2}}{\lambda \lambda^{\prime}}-m^{2} c^{2} \\
& =2 m\left(\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2} /}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \\
\frac{2 h^{2}}{\lambda \lambda^{\prime}}(1-\cos \phi) & =2 m\left(\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)=2 m h c\left(\frac{\lambda^{\prime}-\lambda}{\lambda \lambda^{\prime}}\right)=\frac{2 m h c \Delta \lambda}{\lambda \lambda^{\prime}}
\end{aligned}
$$

## Compton scattering derivation (cont.)

After expansion, cancellation,

$$
\begin{aligned}
\frac{h^{2}}{\lambda^{2}}+\frac{h^{2} /}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \cos \phi & =\left(m c+\frac{h}{\lambda}-\frac{h}{\lambda^{\prime}}\right)^{2}-m^{2} c^{2} \\
& =m^{2} c^{2}+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2}}{\lambda^{\prime 2}}-\frac{2 m c h}{\lambda}-\frac{2 m c h}{\lambda^{\prime}}+\frac{2 h^{2}}{\lambda \lambda^{\prime}}-m^{2} c^{2} \\
& =2 m\left(\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)+\frac{h^{2}}{\lambda^{2}}+\frac{h^{2} /}{\lambda^{\prime 2}}-\frac{2 h^{2}}{\lambda \lambda^{\prime}} \\
\frac{2 h^{2}}{\lambda \lambda^{\prime}}(1-\cos \phi) & =2 m\left(\frac{h c}{\lambda}-\frac{h c}{\lambda^{\prime}}\right)=2 m h c\left(\frac{\lambda^{\prime}-\lambda}{\lambda \lambda^{\prime}}\right)=\frac{2 m h c \Delta \lambda}{\lambda X^{\prime}} \\
\Delta \lambda & =\frac{h}{m c}(1-\cos \phi)
\end{aligned}
$$

## Compton scattering results

Thus, for an electron

## Compton scattering results

Thus, for an electron

$$
\lambda_{c}=\hbar / m c=3.86 \times 10^{-3} \AA
$$

## Compton scattering results

Thus, for an electron

$$
\begin{aligned}
\lambda_{c} & =\hbar / m c=3.86 \times 10^{-3} \AA \\
r_{0} & =\frac{e^{2}}{4 \pi \epsilon_{0} m c^{2}}=2.82 \times 10^{-5} \AA
\end{aligned}
$$

## Compton scattering results

Thus, for an electron

$$
\begin{aligned}
\lambda_{c} & =\hbar / m c=3.86 \times 10^{-3} \AA \\
r_{0} & =\frac{e^{2}}{4 \pi \epsilon_{0} m c^{2}}=2.82 \times 10^{-5} \AA
\end{aligned}
$$

## Compton scattering results

Thus, for an electron

$$
\lambda_{c}=\hbar / m c=3.86 \times 10^{-3} \AA
$$

Comparing the two scattering lengths: $r_{0} / \lambda_{C}=1 / 137$

$$
r_{0}=\frac{e^{2}}{4 \pi \epsilon_{0} m c^{2}}=2.82 \times 10^{-5} \AA
$$



## X-ray absorption



## X-ray absorption



For absorption coefficient $\mu$ and thickness $d z$ the x-ray intensity is attenuated as

## X-ray absorption



For absorption coefficient $\mu$ and thickness $d z$ the x-ray intensity is attenuated as

$$
d l=-l(z) \mu d z
$$

## X-ray absorption



For absorption coefficient $\mu$ and thickness $d z$ the x-ray intensity is attenuated as

$$
d l=-l(z) \mu d z \quad \longrightarrow \frac{d l}{l(z)}=-\mu d z
$$

## X-ray absorption



For absorption coefficient $\mu$ and thickness $d z$ the $x$-ray intensity is attenuated as

$$
d l=-l(z) \mu d z \quad \longrightarrow \frac{d l}{l(z)}=-\mu d z
$$

integrating both sides

## X-ray absorption


integrating both sides
For absorption coefficient $\mu$ and thickness $d z$ the x-ray intensity is attenuated as

$$
\begin{aligned}
d l & =-l(z) \mu d z \quad \longrightarrow \frac{d l}{l(z)}=-\mu d z \\
\int \frac{d l}{l(z)} & =-\int \mu d z
\end{aligned}
$$

## X-ray absorption


integrating both sides
For absorption coefficient $\mu$ and thickness $d z$ the x-ray intensity is attenuated as

$$
\begin{aligned}
d l & =-l(z) \mu d z \longrightarrow \frac{d l}{l(z)}=-\mu d z \\
\int \frac{d l}{l(z)} & =-\int \mu d z \longrightarrow \ln (I)=-\mu z+C
\end{aligned}
$$

## X-ray absorption


integrating both sides
and taking the anti-log
For absorption coefficient $\mu$ and thickness $d z$ the x-ray intensity is attenuated as

$$
\begin{aligned}
d l & =-l(z) \mu d z \longrightarrow \frac{d l}{l(z)}=-\mu d z \\
\int \frac{d l}{l(z)} & =-\int \mu d z \longrightarrow \ln (I)=-\mu z+C
\end{aligned}
$$

## X-ray absorption


integrating both sides
and taking the anti-log
For absorption coefficient $\mu$ and thickness $d z$ the x-ray intensity is attenuated as

$$
\begin{aligned}
d l & =-l(z) \mu d z \longrightarrow \frac{d l}{I(z)}=-\mu d z \\
\int \frac{d l}{I(z)} & =-\int \mu d z \longrightarrow \ln (I)=-\mu z+C
\end{aligned}
$$

$$
I=e^{C} e^{-\mu z}=A e^{-\mu z}
$$

## X-ray absorption


integrating both sides
and taking the anti-log

$$
I=e^{C} e^{-\mu z}=A e^{-\mu z}
$$

if the intensity at $z=0$
is $I_{0}$, then
For absorption coefficient $\mu$ and thickness $d z$ the x-ray intensity is attenuated as

$$
\begin{aligned}
d l & =-l(z) \mu d z \longrightarrow \frac{d l}{I(z)}=-\mu d z \\
\int \frac{d l}{I(z)} & =-\int \mu d z \longrightarrow \ln (I)=-\mu z+C
\end{aligned}
$$

## X-ray absorption


integrating both sides
and taking the anti-log
if the intensity at $z=0$ is $I_{0}$, then

For absorption coefficient $\mu$ and thickness $d z$ the x-ray intensity is attenuated as

$$
\begin{aligned}
d l & =-l(z) \mu d z \longrightarrow \frac{d l}{I(z)}=-\mu d z \\
\int \frac{d l}{I(z)} & =-\int \mu d z \longrightarrow \ln (I)=-\mu z+C
\end{aligned}
$$

$$
I=e^{C} e^{-\mu z}=A e^{-\mu z}
$$

$$
I=I_{0} e^{-\mu z}
$$

## X-ray absorption

integrating both sides
For absorption coefficient $\mu$ and thickness $d z$ the x-ray intensity is attenuated as

$$
\begin{aligned}
d l & =-l(z) \mu d z \longrightarrow \frac{d l}{I(z)}=-\mu d z \\
\int \frac{d l}{l(z)} & =-\int \mu d z \longrightarrow \ln (I)=-\mu z+C
\end{aligned}
$$

and taking the anti-log

$$
I=e^{C} e^{-\mu z}=A e^{-\mu z}
$$

if the intensity at $z=0$

$$
I=I_{0} e^{-\mu z}
$$

This is just Beer's law with an absorption coefficient which depends on x-ray parameters.

## Absorption event



- X-ray is absorbed by an atom


## Absorption event



- X-ray is absorbed by an atom
- Energy is transferred to a core electron


## Absorption event



- X-ray is absorbed by an atom
- Energy is transferred to a core electron
- Electron escapes atomic potential into the continuum


## Absorption event



- X-ray is absorbed by an atom
- Energy is transferred to a core electron
- Electron escapes atomic potential into the continuum
- Ion remains with a core-hole


## Fluorescence emission

An ion with a core-hole is quite unstable $\left(\approx 10^{-15}\right.$ s)


## Fluorescence emission

An ion with a core-hole is quite unstable ( $\approx 10^{-15}$ s)

- After a short time a higher level electron will drop down in energy to fill the core hole

$-\bigcirc$


## Fluorescence emission

An ion with a core-hole is quite unstable $\left(\approx 10^{-15}\right.$ s)

- After a short time a higher level electron will drop down in energy to fill the core hole
- Energy is liberated in the form of a fluorescence photon


## Fluorescence emission

An ion with a core-hole is quite unstable $\left(\approx 10^{-15}\right.$ s)

- After a short time a higher level electron will drop down in energy to fill the core hole
- Energy is liberated in the form of a fluorescence photon
- This leaves a second hole (not core) which is then filled from an even higher shell


## Fluorescence emission

An ion with a core-hole is quite unstable $\left(\approx 10^{-15}\right.$ s)

- After a short time a higher level electron will drop down in energy to fill the core hole
- Energy is liberated in the form of a fluorescence photon
- This leaves a second hole (not core) which is then filled from an even higher shell
- The result is a cascade of fluorescence photons which are characteristic of the absorbing atom


## Auger emission

While fluorescence is the most probable method of core-hole relaxation there are other possible mechanisms


## Auger emission

While fluorescence is the most probable method of core-hole relaxation there are other possible mechanisms

- In the Auger process, a higher level electron will drop down in energy to fill the core hole



## Auger emission

While fluorescence is the most probable method of core-hole relaxation there are other possible mechanisms


- In the Auger process, a higher level electron will drop down in energy to fill the core hole
- The energy liberated causes the secondary emission of an electron


## Auger emission

While fluorescence is the most probable method of core-hole relaxation there are other possible mechanisms


- In the Auger process, a higher level electron will drop down in energy to fill the core hole
- The energy liberated causes the secondary emission of an electron
- This leaves two holes which then filled from higher shells


## Auger emission

While fluorescence is the most probable method of core-hole relaxation there are other possible mechanisms


- In the Auger process, a higher level electron will drop down in energy to fill the core hole
- The energy liberated causes the secondary emission of an electron
- This leaves two holes which then filled from higher shells
- So that the secondary electron is accompanied by fluorescence emissions at lower energies


## Absorption coefficient

The absorption coefficient $\mu$, depends strongly on the x-ray energy $E$, the atomic number of the absorbing atoms $Z$, as well as the density $\rho$, and atomic mass $A$ :
$\mu \sim$

## Absorption coefficient

The absorption coefficient $\mu$, depends strongly on the x-ray energy $E$, the atomic number of the absorbing atoms $Z$, as well as the density $\rho$, and atomic mass $A$ :
$\mu \sim \overline{E^{3}}$

## Absorption coefficient

The absorption coefficient $\mu$, depends strongly on the x-ray energy $E$, the atomic number of the absorbing atoms $Z$, as well as the density $\rho$, and atomic mass $A$ :
$\mu \sim \frac{Z^{4}}{E^{3}}$

## Absorption coefficient

The absorption coefficient $\mu$, depends strongly on the x-ray energy $E$, the atomic number of the absorbing atoms $Z$, as well as the density $\rho$, and atomic mass $A$ :
$\mu \sim \frac{\rho Z^{4}}{E^{3}}$

## Absorption coefficient

The absorption coefficient $\mu$, depends strongly on the x-ray energy $E$, the atomic number of the absorbing atoms $Z$, as well as the density $\rho$, and atomic mass $A$ :
$\mu \sim \frac{\rho Z^{4}}{A E^{3}}$

## Absorption coefficient

The absorption coefficient $\mu$, depends strongly on the x-ray energy $E$, the atomic number of the absorbing atoms $Z$, as well as the density $\rho$, and atomic mass $A$ :


## Absorption coefficient

Isolated gas atoms show a sharp jump and a smooth curve



## Absorption coefficient

Isolated gas atoms show a sharp jump and a smooth curve Atoms in a solid or liquid show fine structure after the absorption edge called XANES and EXAFS



## Absorption coefficient

For an elemental material, the absorption coefficient, $\mu$, is simply the product of the atomic density, $\rho_{a}$, times the atomic absorption cross-section, $\sigma_{a}$.

## Absorption coefficient

For an elemental material, the absorption coefficient, $\mu$, is simply the product of the atomic density, $\rho_{a}$, times the atomic absorption cross-section, $\sigma_{a}$.

$$
\mu=\rho_{\mathrm{a}} \sigma_{a}
$$

## Absorption coefficient

For an elemental material, the absorption coefficient, $\mu$, is simply the product of the atomic density, $\rho_{a}$, times the atomic absorption cross-section, $\sigma_{a}$.
rewriting in terms of the mass density, $\rho_{m}$, the atomic mass, $M_{a}$, and Avogadro's number, $N_{A}$, the absorption coefficient becomes

$$
\mu=\rho_{a} \sigma_{a}
$$

## Absorption coefficient

For an elemental material, the absorption coefficient, $\mu$, is simply the product of the atomic density, $\rho_{a}$, times the atomic absorption cross-section, $\sigma_{a}$.
rewriting in terms of the mass density, $\rho_{m}$, the atomic mass, $M_{a}$, and Avogadro's number, $N_{A}$, the absorption coefficient becomes

$$
\mu=\rho_{a} \sigma_{a}=\left(\frac{\rho_{m} N_{A}}{M_{a}}\right) \sigma_{a}
$$

## Absorption coefficient

For an elemental material, the absorption coefficient, $\mu$, is simply the product of the atomic density, $\rho_{\mathrm{a}}$, times the atomic absorption cross-section, $\sigma_{a}$.
rewriting in terms of the mass density, $\rho_{m}$, the atomic mass, $M_{a}$, and Avogadro's number, $N_{A}$, the absorption coefficient becomes

$$
\mu=\rho_{a} \sigma_{a}=\left(\frac{\rho_{m} N_{A}}{M_{a}}\right) \sigma_{a}
$$

if the absorber is made up of a number of different atoms, this calculation can be generalized

## Absorption coefficient

For an elemental material, the absorption coefficient, $\mu$, is simply the product of the atomic density, $\rho_{a}$, times the atomic absorption cross-section, $\sigma_{a}$.
rewriting in terms of the mass density, $\rho_{m}$, the atomic mass, $M_{a}$, and Avogadro's number, $N_{A}$, the absorption coefficient becomes

$$
\mu=\rho_{a} \sigma_{a}=\left(\frac{\rho_{m} N_{A}}{M_{a}}\right) \sigma_{a}
$$

if the absorber is made up of a number of different atoms, this calculation can be generalized

$$
\mu=\sum_{j} \rho_{j} \sigma_{a j}
$$

## Absorption coefficient

For an elemental material, the absorption coefficient, $\mu$, is simply the product of the atomic density, $\rho_{\mathrm{a}}$, times the atomic absorption cross-section, $\sigma_{a}$.
rewriting in terms of the mass density, $\rho_{m}$, the atomic mass, $M_{a}$, and Avogadro's number, $N_{A}$, the absorption coefficient becomes

$$
\mu=\rho_{\mathrm{a}} \sigma_{a}=\left(\frac{\rho_{m} N_{A}}{M_{a}}\right) \sigma_{a}
$$

if the absorber is made up of a number of different atoms, this calculation can be generalized

$$
\mu=\sum_{j} \rho_{j} \sigma_{a j}
$$

where $\rho_{j}$ and $\sigma_{a j}$ are the atomic density and atomic absorption cross-section of each component

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass absorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab- $\quad \mu_{m}=\mu / \rho$ sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab- $\quad \mu_{m}=\mu / \rho$ sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$

Beer's Law now becomes

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab- $\quad \mu_{m}=\mu / \rho$ sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$

Beer's Law now becomes

$$
I=I_{0} e^{-\mu z}=I_{0} e^{-\mu_{m} \rho z}
$$

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab- $\quad \mu_{m}=\mu / \rho$ sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$

Beer's Law now becomes $\quad I=I_{0} e^{-\mu z}=I_{0} e^{-\mu_{m} \rho z}$
Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area $A$

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab- $\quad \mu_{m}=\mu / \rho$ sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$

Beer's Law now becomes

$$
I=I_{0} e^{-\mu z}=I_{0} e^{-\mu_{m} \rho z}
$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area $A$

If the compound is made up of $x_{j}$ atoms with atomic mass $M_{j}$ and has a molecular mass $M_{c}$ and density $\rho_{c}$, we can write:

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab- $\quad \mu_{m}=\mu / \rho$ sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$

Beer's Law now becomes

$$
I=I_{0} e^{-\mu z}=I_{0} e^{-\mu_{m} \rho z}
$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area $A$

If the compound is made up of $x_{j}$ atoms with atomic mass $M_{j}$ and has a molecular mass

$$
M_{c}=\sum_{j} x_{j} M_{j}
$$ $M_{c}$ and density $\rho_{c}$, we can write:

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab- $\quad \mu_{m}=\mu / \rho$ sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$

Beer's Law now becomes $\quad I=I_{0} e^{-\mu z}=I_{0} e^{-\mu_{m} \rho z}$
Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area $A$

If the compound is made up of $x_{j}$ atoms with atomic mass $M_{j}$ and has a molecular mass $M_{c}$ and density $\rho_{c}$, we can write:

$$
\begin{aligned}
M_{c} & =\sum_{j} x_{j} M_{j} \\
\mu_{m} & =\left(N_{A} / M_{c}\right) \sum_{j} x_{j} \sigma_{a j}
\end{aligned}
$$

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab-

$$
\mu_{m}=\mu / \rho
$$

sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$
Beer's Law now becomes

$$
I=I_{0} e^{-\mu z}=I_{0} e^{-\mu_{m} \rho z}
$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area $A$

If the compound is made up of $x_{j}$ atoms with atomic mass $M_{j}$ and has a molecular mass $M_{c}$ and density $\rho_{c}$, we can write:

The "thickness" of a mass $m$ of the compound, distributed over an area $A$ is then:

$$
\begin{aligned}
M_{c} & =\sum_{j} x_{j} M_{j} \\
\mu_{m} & =\left(N_{A} / M_{c}\right) \sum_{j} x_{j} \sigma_{a j}
\end{aligned}
$$

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab-

$$
\mu_{m}=\mu / \rho
$$

sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$
Beer's Law now becomes

$$
I=I_{0} e^{-\mu z}=I_{0} e^{-\mu_{m} \rho z}
$$

Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area $A$

If the compound is made up of $x_{j}$ atoms with atomic mass $M_{j}$ and has a molecular mass $M_{c}$ and density $\rho_{c}$, we can write:

The "thickness" of a mass $m$ of the compound, distributed over an area $A$ is then:

$$
\begin{aligned}
M_{c} & =\sum_{j} x_{j} M_{j} \\
\mu_{m} & =\left(N_{A} / M_{c}\right) \sum_{j} x_{j} \sigma_{a j}
\end{aligned}
$$

$$
z=\frac{m}{\rho_{c} A}
$$

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab-

$$
\mu_{m}=\mu / \rho
$$

sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$
Beer's Law now becomes $\quad I=I_{0} e^{-\mu z}=I_{0} e^{-\mu_{m} \rho z}$
Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area $A$

If the compound is made up of $x_{j}$ atoms with atomic mass $M_{j}$ and has a molecular mass $M_{c}$ and density $\rho_{c}$, we can write:

The "thickness" of a mass $m$ of the compound, distributed over an area $A$ is then:

$$
\begin{aligned}
M_{c} & =\sum_{j} x_{j} M_{j} \\
\mu_{m} & =\left(N_{A} / M_{c}\right) \sum_{j} x_{j} \sigma_{a j}
\end{aligned}
$$

$$
z=\frac{m}{\rho_{c} A}
$$

This leads to an absorption per unit mass of $\mu_{m} / A$ and Beer's law becomes

## Absorption coefficient of a compound

$\mu\left[\mathrm{cm}^{-} 1\right]$ is the linear absorption coefficient. It is useful in practice to define the mass ab-

$$
\mu_{m}=\mu / \rho
$$

sorption coefficient, $\mu_{m}\left[\mathrm{~cm}^{2} / \mathrm{g}\right]$
Beer's Law now becomes $\quad I=I_{0} e^{-\mu z}=I_{0} e^{-\mu_{m} \rho z}$
Suppose we want to compute the absorption coefficient per unit mass of a compound if we distribute it over an area $A$

If the compound is made up of $x_{j}$ atoms with atomic mass $M_{j}$ and has a molecular mass $M_{c}$ and density $\rho_{c}$, we can write:

The "thickness" of a mass $m$ of the com-
pound, distributed over an area $A$ is then:
This leads to an absorption per unit mass of
pound, distributed over an area $A$ is then:
This leads to an absorption per unit mass of

$$
\begin{aligned}
& M_{c}=\sum_{j} x_{j} M_{j} \\
& \mu_{m}=\left(N_{A} / M_{c}\right) \sum_{j} x_{j} \sigma_{a j}
\end{aligned}
$$

$$
z=\frac{m}{\rho_{c} A}
$$

$\mu_{m} / A$ and Beer's law becomes

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have
$I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of
$\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\quad \rho=5.24 \mathrm{~g} / \mathrm{cm}^{2}$ $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\quad \rho=5.24 \mathrm{~g} / \mathrm{cm}^{2}$ $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

$$
M_{F e}=55.895 \mathrm{~g} / \mathrm{mol}
$$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{F e} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol}
\end{aligned}
$$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{\mathrm{Fe}} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{\mathrm{O}} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{c} & =159.69 \mathrm{~g} / \mathrm{mol}
\end{aligned}
$$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{\mathrm{Fe}} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{c} & =159.69 \mathrm{~g} / \mathrm{mol}
\end{aligned}
$$

begin by finding tabulated values of the cross-section for the elements Fe and O at 5 keV

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{F e} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{c} & =159.69 \mathrm{~g} / \mathrm{mol} \\
\sigma_{F e} & =138.860 \mathrm{~cm}^{2} / \mathrm{g}
\end{aligned}
$$

begin by finding tabulated values of the cross-section for the elements Fe and O at 5 keV

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{F e} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{c} & =159.69 \mathrm{~g} / \mathrm{mol} \\
\sigma_{\mathrm{Fe}} & =138.860 \mathrm{~cm}^{2} / \mathrm{g} \\
\sigma_{O} & =46.666 \mathrm{~cm}^{2} / \mathrm{g}
\end{aligned}
$$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are
begin by finding tabulated values of the cross-section for the elements Fe and O at 5 keV

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{F e} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{c} & =159.69 \mathrm{~g} / \mathrm{mol} \\
\sigma_{F e} & =138.860 \mathrm{~cm}^{2} / \mathrm{g} \\
\sigma_{O} & =46.666 \mathrm{~cm}^{2} / \mathrm{g}
\end{aligned}
$$

assuming a 5 mm diameter pellet

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{F e} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{\mathrm{c}} & =159.69 \mathrm{~g} / \mathrm{mol} \\
\sigma_{F e} & =138.860 \mathrm{~cm}^{2} / \mathrm{g} \\
\sigma_{O} & =46.666 \mathrm{~cm}^{2} / \mathrm{g} \\
A & =\pi\left(0.25 \mathrm{~cm}^{2}=0.1963 \mathrm{~cm}^{2}\right.
\end{aligned}
$$

begin by finding tabulated values of

$$
\text { the cross-section for the elements } \mathrm{Fe}
$$ and O at 5 keV

assuming a 5 mm diameter pellet

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{F e} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{c} & =159.69 \mathrm{~g} / \mathrm{mol} \\
\sigma_{F e} & =138.860 \mathrm{~cm}^{2} / \mathrm{g} \\
\sigma_{O} & =46.666 \mathrm{~cm}^{2} / \mathrm{g} \\
A & =\pi(0.25 \mathrm{~cm})^{2}=0.1963 \mathrm{~cm}^{2}
\end{aligned}
$$

begin by finding tabulated values of the cross-section for the elements Fe and O at 5 keV
assuming a 5 mm diameter pellet

$$
\mu_{m}=\frac{1}{159.69}[2 \cdot 55.895 \cdot 138.860+3 \cdot 16.000 \cdot 46.666]
$$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{F e} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{c} & =159.69 \mathrm{~g} / \mathrm{mol} \\
\sigma_{F e} & =138.860 \mathrm{~cm}^{2} / \mathrm{g} \\
\sigma_{O} & =46.666 \mathrm{~cm}^{2} / \mathrm{g} \\
A & =\pi(0.25 \mathrm{~cm})^{2}=0.1963 \mathrm{~cm}^{2}
\end{aligned}
$$

begin by finding tabulated values of the cross-section for the elements Fe and O at 5 keV
assuming a 5 mm diameter pellet

$$
\mu_{m}=\frac{1}{159.69}[2 \cdot 55.895 \cdot 138.860+3 \cdot 16.000 \cdot 46.666]=111.23 \mathrm{~cm}^{2} / \mathrm{g}
$$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{F e} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{c} & =159.69 \mathrm{~g} / \mathrm{mol} \\
\sigma_{F e} & =138.860 \mathrm{~cm}^{2} / \mathrm{g} \\
\sigma_{O} & =46.666 \mathrm{~cm}^{2} / \mathrm{g} \\
A & =\pi\left(0.25 \mathrm{~cm}^{2}=0.1963 \mathrm{~cm}^{2}\right.
\end{aligned}
$$

begin by finding tabulated values of the cross-section for the elements Fe and O at 5 keV
assuming a 5 mm diameter pellet

$$
\mu_{m}=\frac{1}{159.69}[2 \cdot 55.895 \cdot 138.860+3 \cdot 16.000 \cdot 46.666]=111.23 \mathrm{~cm}^{2} / \mathrm{g}
$$

$$
\mu_{m} / A=566.7 \mathrm{~g}^{-1}
$$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are
begin by finding tabulated values of

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{F e} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{\mathrm{c}} & =159.69 \mathrm{~g} / \mathrm{mol} \\
\sigma_{F e} & =138.860 \mathrm{~cm}^{2} / \mathrm{g} \\
\sigma_{O} & =46.666 \mathrm{~cm}^{2} / \mathrm{g} \\
A & =\pi\left(0.25 \mathrm{~cm}^{2}=0.1963 \mathrm{~cm}^{2}\right.
\end{aligned}
$$ the cross-section for the elements Fe and O at 5 keV

assuming a 5 mm diameter pellet

$$
\mu_{m}=\frac{1}{159.69}[2 \cdot 55.895 \cdot 138.860+3 \cdot 16.000 \cdot 46.666]=111.23 \mathrm{~cm}^{2} / \mathrm{g}
$$

$$
\mu_{m} / A=566.7 \mathrm{~g}^{-1} \quad \mu=\mu_{\mathrm{m}} \rho=582.9 \mathrm{~cm}^{-1}
$$

## Absorption of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ at 5 keV

The most commonly tabulated cross-sections are not the atomic cross-sections but the mass cross sections, $\sigma_{j}=N_{A} \sigma_{a j} / M_{j}$ so we have $I=I_{0} e^{-\left(\mu_{m} / A\right) m}, \mu_{m}=\frac{N_{A}}{M_{c}} \sum_{j} x_{j} \sigma_{a j}=\frac{N_{A}}{M_{c}} \sum_{j} \frac{M_{j}}{N_{A}} x_{j} \sigma_{j}=\frac{1}{M_{c}} \sum_{j} M_{j} x_{j} \sigma_{j}$
the molecular mass and density of $\mathrm{Fe}_{2} \mathrm{O}_{3}$ are
begin by finding tabulated values of

$$
\begin{aligned}
\rho & =5.24 \mathrm{~g} / \mathrm{cm}^{2} \\
M_{F e} & =55.895 \mathrm{~g} / \mathrm{mol} \\
M_{O} & =16.000 \mathrm{~g} / \mathrm{mol} \\
M_{\mathrm{c}} & =159.69 \mathrm{~g} / \mathrm{mol} \\
\sigma_{F e} & =138.860 \mathrm{~cm}^{2} / \mathrm{g} \\
\sigma_{O} & =46.666 \mathrm{~cm}^{2} / \mathrm{g} \\
A & =\pi\left(0.25 \mathrm{~cm}^{2}=0.1963 \mathrm{~cm}^{2}\right.
\end{aligned}
$$ the cross-section for the elements Fe and O at 5 keV

assuming a 5 mm diameter pellet

$$
\mu_{m}=\frac{1}{159.69}[2 \cdot 55.895 \cdot 138.860+3 \cdot 16.000 \cdot 46.666]=111.23 \mathrm{~cm}^{2} / \mathrm{g}
$$

$$
\mu_{\mathrm{m}} / A=566.7 \mathrm{~g}^{-1} \quad \mu=\mu_{\mathrm{m}} \rho=582.9 \mathrm{~cm}^{-1} \quad 1 / \mu=17.2 \mu \mathrm{~m}
$$

## Comparison of cross sections



## Comparison of cross sections




## Comparison of cross sections




Photoelectric absorption dominates at low energies

## Comparison of cross sections




Photoelectric absorption dominates at low energies
Thomson scattering (coherent) drops rapidly with energy

## Comparison of cross sections




Photoelectric absorption dominates at low energies
Thomson scattering (coherent) drops rapidly with energy
Compton scattering (incoherent) dominates at medium energies

## Comparison of cross sections




Photoelectric absorption dominates at low energies
Thomson scattering (coherent) drops rapidly with energy
Compton scattering (incoherent) dominates at medium energies Pair production dominates at high energies

